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**ENKELEDA DERVISHI ET AL: "Versatile Catalytic System for the Large-Scale and Controlled Synthesis of Single-Wall, Double-Wall, Multi-Wall, and Graphene Carbon Nanostructures", CHEMISTRY OF MATERIALS, vol. 21, no. 22, 24 November 2009 (2009-11-24), pages 5491-5498, XP55040031, ISSN: 0897-4756, DOI: 10.1021/cm902502c**  
**JIN X ET AL: "Understanding the kinetics of coal electrolysis at intermediate temperatures", JOURNAL OF POWER SOURCES, ELSEVIER SA, CH, vol. 195, no. 15, 1 August 2010 (2010-08-01), pages 4935-4942, XP026991263, ISSN: 0378-7753 [retrieved on 2010-02-10]**



# DESCRIPTION

## BACKGROUND OF THE INVENTION

**[0001]** Graphene, a two dimensional monolayer of  $sp^2$ -hybridized carbon atoms arranged in a honeycomb network, exhibits excellent mechanical, thermal, electrical, and optical properties; large specific surface area; and chemical stability. These superb properties offer graphenes many potential applications ranging from nanoelectronics, composite materials, sensors to electrochemical electrodes in lithium ion batteries, solar cells, or ultracapacitors. Specific electronics applications of graphene include mobile phones, radars, transistors, data storage, touch screens, and wearable electronics.

**[0002]** Although there is an increasing interest in both the theoretical and experimental study of graphene, high quality, low cost, and scalable graphene production is still a major challenge. There are basically four methods of making graphene. The first method is the mechanical exfoliation of graphite, which can produce high quality graphene. However, the yield is low because it is hard to control the layers of graphene in the product. Therefore, this method is only useful for fundamental study. The second is chemically derived graphene colloidal suspensions. This method can produce scalable graphene sheets in solution, but the electrical conductivity is poor since the chemical oxidation and reduction process may lead to structural defects. The third method is organic synthesis. However, graphene produced by total organic synthesis has size limits due to insolubility of macromolecules and side reactions. Lastly, chemical vapor deposition (CVD) is a promising method for growing high quality and large-scale graphene. The existing CVD methods require expensive transition metals such as Co, Ni, Pt, Ir, and Ru as catalyst, hydrocarbon gas as carbon source, and e-beam evaporation process. Some CVD methods also need ultrahigh vacuum conditions and/or specific substrates for graphene growth. High costs hinder the use of the CVD methods for large-scale graphene production and applications.

**[0003]** The paper "Versatile Catalytic System for the Large-Scale and Controlled Synthesis of Single-Wall, Double-Wall, Multi-Wall, and Graphene Carbon Nanostructures", CHEMISTRY OF MATERIALS, vol. 21, no. 22, 24 November 2009 (2009-11-24), pages 5491-5498, discloses a method of forming carbon nanostructures through radio-frequency catalytic chemical vapour deposition.

## SUMMARY OF THE INVENTION

**[0004]** The invention is defined by the claims.

**[0005]** The present invention is premised on the realization that graphene can be inexpensively produced using a coal byproduct. More particularly, the present invention is

premised on the realization that char, which is the byproduct of electrolysis of an aqueous coal slurry, can be used to form graphene.

**[0006]** More particularly, ground coal can be subjected to electrolysis which will consume a portion of the coal, forming particles that are coated with a gelatinous organometallic material which prevents further electrolysis. These coated particles can be subjected to solvent extraction to remove the organic material, allowing the coal particles to be used again for electrolysis. This can be repeated until cleaning or solvent extraction of the coating no longer improves the efficacy of the coal particles in electrolysis.

**[0007]** The char, electrolyzed coal particles or the gelatinous coating can be used as a carbon source material to form graphene. This method can be used with a wide variety of different coals, and, further, the electrolyzed coal can be collected from the electrolysis process at any time, either after an initial electrolysis or after repeated electrolysis, depending on the end use requirements. Graphene is produced from electrolyzed coal by chemical vapor deposition at atmospheric pressure in the presence of hydrogen gas. This produces high quality graphene at a relatively low cost and without expensive catalysts.

**[0008]** The objects and advantages of the present invention will be further appreciated in light of the following detailed description and examples in which:

#### **BRIEF DESCRIPTION OF THE DRAWING**

##### **[0009]**

FIG. 1 is a diagrammatic depiction of an electrolytic apparatus to form hydrogen from coal slurries;

FIG. 2 is an exploded diagrammatic view of the electrolytic cell of FIG. 1; and

FIG. 3 is a diagrammatic depiction of an apparatus used to form graphene.

#### **DETAILED DESCRIPTION**

**[0010]** FIG. 1 shows a diagrammatic depiction of an apparatus 10 used to electrolyze coal which, in turn, produces char or electrolyzed coal particles for the manufacture of graphene. As shown in FIG. 1, the apparatus 10 includes an electrolytic cell 11, which incorporates a cathode 12 and an anode 14 separated by a membrane 13, such as a Nafion membrane or polyethylene membrane.

**[0011]** A reservoir 16, which contains the anode solution, leads to a pump 17, which pumps the

anode solution into the anodic side of the electrolytic cell 11. As shown, the anode solution flows through line 23 through the channels 21 in an acrylic block 22. The anode fluid passes through the channels 21 and returns via line 24 to the reservoir 16. The generated gas, carbon dioxide, in the fluid is emitted from the reservoir and is directed to a gas collector 26. A temperature controller 27 is located in the reservoir 16.

**[0012]** At the opposite side, reservoir 30 includes the cathode solution, which is directed to pump 32 leading to the cathode side of electrolytic cell 11. Again, this cathode solution passes through line 33 through channels (not shown) in the acrylic block 34, which provide contact with the cathode 12, and the fluid then is directed from the channels through line 36 back to the reservoir 30. The generated gas, hydrogen, is directed to the gas collector 40. Again, a temperature controller 42 is located in reservoir 30.

**[0013]** FIG. 2 shows an exploded view of the electrolytic cell 11. Acrylic blocks 22 and 34 are mirror images of each other. The anode 14 and cathode 12 are on either side of the separator 13. Spacing is established between the acrylic blocks 22 and 34, the cathode 12, and the anode 14 by inert separators 44 and 46. These can be, for example, polytetrafluoroethylene, Nafion or polyethylene.

**[0014]** The cathode 12 can be any material which will withstand the acidic conditions in the electrolytic cell 11. These include carbon, nickel, and noble metals such as platinum, iridium, rhodium, and combinations thereof.

**[0015]** Although the anode 14 can be any conductor, the anode 14 typically comprises a noble metal-containing electrocatalyst electroplated, i.e., deposited, on a substrate or support. The electrocatalyst may be single metal, bi- or trimetallic, and comprise at least one noble metal and one or more other metals that are active to coal electrolysis. The other metals may be, but are not necessarily, noble metals. In some embodiments, the electrocatalyst may comprise a single noble metal on a support.

**[0016]** The support may be chosen from many known supports. Some suitable supports include noble metal meshes and foils, such as platinum mesh, platinum foil, titanium mesh, Hastelloy mesh, gold mesh, gold foil, tantalum mesh, tantalum foil, as well as platinum or iridium sponges. When mesh is used as the substrate, the mesh size will be chosen such that it can be properly electroplated with the electrocatalyst, whether it is a bi- or tri-metallic catalyst electroplated on the substrate, or a bi- or trimetallic/Raney metal catalyst electro deposited on the substrate. Aside from the specific substrates listed, other suitable supports will be recognized by those of ordinary skill in the art. In some embodiments, the electrode is a bi- or tri-metallic electrocatalyst electroplated directly on a support.

**[0017]** Suitable metals for bi- and tri-metallic catalysts are selected from platinum, iridium, ruthenium, rhenium, palladium, gold, silver, nickel, and iron. By way of example, in one embodiment, the electrode is a platinum-iridium electrocatalyst electrodeposited on platinum mesh.

**[0018]** The electrolysis of coal is further described in U.S. Patent No. 7,736,475. The Application references coal as a fuel. Coal is intended to reference a carbon based product such a coal, charcoal lignites, or graphite.

**[0019]** To form char or electrolyzed coal, a coal slurry is prepared by grinding coal to a particle size between 1 $\mu$ m to 249  $\mu$ m. The pulverized coal is mixed with an electrolyte containing a proton carrier (e.g., H<sub>2</sub>SO<sub>4</sub>, H<sub>3</sub>PO<sub>4</sub>, etc.) and a catalytic salt (e.g., iron salt, or cerium salt). Typically, iron is required for an efficient reaction. The concentration of iron can be up to 10,000 ppm.

**[0020]** The current density which can be applied in the electrolytic process can be from 30 mA/cm<sup>2</sup> up to 200 mA/cm<sup>2</sup>. It is preferred to use a current density of at least 90 mA/ cm<sup>2</sup> or higher. The temperature which can be applied in the electrolytic process can be from 80° C or higher. The coal slurry is introduced into the coal electrolytic cell as the anode solution. During the electrolysis, dehydrogenation of the coal takes place, producing pure hydrogen at the cathodic compartment of the electrolytic cell. Coal is oxidized to carbon dioxide (which is collected at the anodic compartment of the electrolytic cell) and large hydrocarbon structures, which coat the coal particles. After electrolysis, the electrolyzed char slurry is filtered to separate the electrolyte from the solids. The electrolyte and catalytic salts can be reused in the coal electrolytic cell.

**[0021]** The mechanism by which coal is electrolyzed is a multistep process. Fe (III) ions in solution are adsorbed on the surface of a coal particle. The adsorbed C-Fe (III)<sub>ads</sub> structure flows toward the surface of the electrode due to forced flow. The C-Fe (III)<sub>ads</sub> structure contacts the anode electrode, and Fe (III)<sub>ads</sub> on the coal acts as a bridge between the coal and the electrode. Due to steric effects and/or electrostatic charges, Fe (III) is de-adsorbed from the coal particle to the anode electrode, and then the oxidation of coal takes place, simultaneously reducing Fe (III) to Fe (II). During the process coal can be oxidized to CO<sub>2</sub> and/or other large chain hydrocarbons. The Fe (II) ions are oxidized at the anode of the electrolytic cell to regenerate Fe (III) ions. As coal oxidizes, gelatinous films grow on the surface of the coal particle, which ultimately prevent the oxidation of coal.

**[0022]** Thus, the coal slurry is passed through the apparatus 10 shown in FIG. 1 and repeatedly circulated back to the reservoir 16 and through the electrolytic cell 11. Eventually, the coal particles become coated with the gelatinous material, which prevents their further use in electrolysis. These particles can be then directed via line 44 to a solvent extractor 48 which contacts the coal particles with an organic solvent such as ethanol, isopropanol, pyridine, or acetone.

**[0023]** The solvent at elevated temperatures, if necessary, removes the gelatinous coating from the coal particles. The solvent is separated from the particles by filtration, and the particles can then be introduced back into the anode solution through line 49 and electrolyzed

further. These steps can be repeated until the coal particles are no longer electrolyzed (i.e., until, the production of hydrogen is significantly reduced). At this point, the coal particles coated with the gelatinous coating can be used in the formation of graphene. Coal particles can be used for the formation of graphene at any cycle during the electrolysis; however, it is preferable to use this feedstock for graphene production when the hydrogen production significantly decreases.

**[0024]** Once it is determined that the coal particles should be used for formation of graphene, the coal particles are filtered to separate the electrolyte, which can then be reused. The particles with the gelatinous coating are then dried.

**[0025]** The graphene can be formed from the dried char, that is, the electrolyzed coal particles with the gelatinous film, from the gelatinous film, or from the particles without the gelatinous film. If formed from the particles coated with the gelatinous film, the char is simply separated from the electrolyte, as described above. This electrolyzed coal is then dried to evaporate the water, as described above, and coated onto a support or substrate 50, such as a copper film. The electrolyzed coal can be dispersed in a carrier such as liquid alcohols, alkenes, or tones, and sprayed onto the substrate 50. In an embodiment, a gas sprayer may be used to dispense the electrolyzed coal onto the substrate 50. The carrier is then evaporated. Alternatively, the dried, ground particles can be applied as a thin coating onto the substrate 50. It is preferable to obtain a very thin coating of the electrolyzed coal in the substrate 50. Using the carrier to disperse the electrolyzed coal on the substrate facilitates this.

**[0026]** In a second method, the gelatinous material, which has been separated from the electrolyzed coal particles, is concentrated by removing extraction solvents. This produces a thick carbonaceous syrup, which is then coated onto the substrate 50. This can be applied, again, with a carrier, which is subsequently evaporated, or directly applied onto the substrate 50.

**[0027]** In a third method, the electrolyzed coal particles separated from the gelatinous material can also be used as the carbon source for the formation of graphene. When utilizing this embodiment, the electrolyzed coal is separated from the electrolyte and dried. The gelatinous material is removed, leaving the dried electrolyzed coal particles. These particles may be dispersed in a carrier and sonicated to disperse the particles in the solution. The particles can then be sprayed onto the substrate 50 and dried. Alternatively, the dried, ground, particles can be applied as a thin coating onto the substrate 50. Again, with all of these embodiments it is preferable to have as thin a coating as possible on the substrate 50 in order to facilitate production of graphene.

**[0028]** Prior to deposition of the electrolyzed coal, the substrate 50 can be rinsed with acetone or isopropanol (IPA) and sonicated for a period of time, for example, 10-15 minutes. Then, the substrate 50 may further be pre-heated in a furnace to a temperature of about 800° C for 30 minutes.

**[0029]** FIG. 3 shows a diagrammatic depiction of the apparatus 60 used to produce graphene. As shown, this apparatus 60 includes a gas source 62. The gas source 62 may include, for example, hydrogen and an inert carrier such as argon. The gas is introduced into a quartz tube heating element 64, which contains the substrate 50, such as copper foils, coated with the electrolyzed coal product. Also located in the quartz tube 64 is an uncoated substrate or surface 66 on which the graphene will be produced. The surface 66 for supporting nanostructure synthesis may be a copper foil, a silicon wafer, or any other material that can withstand reaction conditions. The quartz tube 64 is inserted into the heating zone of a tube furnace 68 where the graphene formation takes place.

**[0030]** The graphene is then formed by flowing a reductant gas over the coated substrate 50 at elevated temperatures using chemical vapor deposition. Any commercially available chemical vapor deposition apparatus is suitable, and the graphene formation can be conducted at atmospheric pressure.

**[0031]** Higher temperatures facilitate the production of graphene. At lower temperatures, such as about 400° C, a carbon film is produced which is not graphene. Accordingly, it is preferred to use higher temperatures, at least 490° C or higher, up to about 1100° C. Higher temperatures, such as 700 to 1000° C can also be used, and, typically, 800 to 1000° C will be utilized. These temperatures cause graphite production in the presence of reductant gas. A flowing stream of reductant gas deposits graphene onto the surface 66.

**[0032]** With the electrolyzed coal on substrate 50, an inert gas such as argon is introduced through the apparatus 60 at a rate of, for example, 600 SCCM to purge the system of any reactive gases. The duration of this purge may last, for example, 20 minutes. Then, over a period of, for example, 40 minutes, the temperature in the furnace is increased to reaction conditions, typically 800° C. A reductant gas, in particular hydrogen, is then introduced into the system along with the argon, typically at 600 SCCM of argon with 100 SCCM of hydrogen. The exact ratio of the gases is not critical. During this time, graphene growth occurs. The graphene growth period can be continued as long as necessary. Typically, the time can be from a few minutes to an hour or more. As an example, the graphene growth period may have a duration of 30 minutes. When desired, the reaction is stopped by discontinuing the introduction of the hydrogen and cooling the furnace 68 rapidly to room temperature. The manufacture of graphene from electrolyzed coal will be further appreciated in light of the following examples, 1-5.

#### **Example 1**

**[0033]** In one example, a slurry having coal particles of approximately 44  $\mu\text{m}$  were suspended in a concentration of approximately 0.04 g/ml. The slurry had an iron concentration of approximately 40 mM. The coal slurry was electrolyzed at a temperature of 104° C and with a current of 100 mA/cm<sup>2</sup>.



**[0034]** After the electrolysis cycle was repeated several times, the electrolyzed coal particles were used for graphene synthesis. A copper foil having dimensions of approximately 2 cm x 1 cm was pre-treated with an acetone rinse. Approximately 10 mg of electrolyzed coal was then deposited on the copper foil. The copper foil having the electrolyzed coal deposited thereon, as well as a clean copper foil (i.e., having no electrolyzed coal deposited thereon) were positioned inside a ¼ in. diameter quartz tube. The quartz tube was then moved inside of a furnace.

**[0035]** 600 SCCM argon flowed over the wafer for 20 minutes to create an inert atmosphere. The furnace was then ramped up to 1000° C in 40 minutes with a 600 SCCM argon and 100 SCCM hydrogen flow. The 1000° C temperature was maintained for 30 minutes, during which graphene growth occurred. The wafer was then cooled down by moving out of the furnace. The wafer initially experienced rapid cool down and was fully cooled to room temperature over a period of 3 hours, during which time the argon-hydrogen flow continued to flow over the wafer.

**[0036]** Following the graphene growth cycle and cool down, the copper foil on which the electrolyzed coal had been deposited was placed in Marble's reagent, which dissolved the copper foil. At this point, a graphene film was left floating in the reagent, and the film was then transferred to a titanium foil. Raman spectroscopy of the graphene film showed significantly fewer defects than regular coal. A selected area electron diffraction (SAED) pattern showed a hexagonal lattice structure of the single crystal graphene.

### **Example 2**

**[0037]** In another example, a silicon wafer was pre-treated, and IPA was deposited on the wafer rather than electrolyzed coal. The wafer was placed in the furnace, and argon was purged at 1 SLPM for 20 minutes. The furnace was then ramped up to 800° C in 30 minutes with an argon flow. The 800° C temperature was maintained for 30 minutes with a hydrogen-nitrogen flow at 0.8 SLPM. The system was then cooled down to room temperature under 1 SLPM of argon flow.

**[0038]** In this example, no graphene synthesized on the wafer. This example suggests that IPA by itself does not contribute to the graphene synthesis.

### **Example 3**

**[0039]** In yet another example, a slurry having coal particles of approximately 210-249 µm suspended in a concentration of approximately 0.04 g/ml was made. The slurry had an iron concentration of approximately 40 mM. The coal slurry was electrolyzed at a temperature of 104° C and with a current of 100 mA/cm<sup>2</sup>. The resulting electrolyzed coal was then used for graphene synthesis.

[0040] For one substrate, approximately 6 mg of the electrolyzed coal was diluted in 6 ml IPA. The solution was sonicated for 5 minutes. The solution was then diluted 30 times using IPA and sonicated again. 25 ml of the solution was deposited on a silicon wafer, and the IPA was allowed to evaporate. This same pre-treatment process was also used to prepare a second silicon wafer substrate on which raw coal particles in the range of 210-249  $\mu\text{m}$  were deposited.

[0041] In separate reactions, the wafer with electrolyzed coal and the wafer with raw coal were moved into the furnace. Each wafer was purged with argon at 1 SLPM for 20 minutes. For each wafer, the furnace was ramped up to 800° C with a flow of a hydrogen-nitrogen mixture at 0.8 SLPM for 30 minutes. The system was then cooled down to room temperature under an argon flow of 1 SLPM.

[0042] A comparison of the two treated wafers showed that graphene synthesized on the wafer with electrolyzed coal. On the other hand, while nanostructures grew on the wafer having raw coal, the nanostructure topography was different than the expected topography for graphene. Therefore, this example suggests that raw coal alone by itself does not contribute to graphene synthesis.

#### Example 4

[0043] In yet another example, synthesis was performed with samples of coal that were electrolyzed in different conditions, as well as with raw coal. Those conditions are summarized in Table 1.

Table 1

	Sample A	Sample B	Sample C	Raw Coal
Temperature (° C)	80	104	104	--
Particle size ( $\mu\text{m}$ )	44	44	44	44
Coal concentration (g/ml)	0.04	0.04	0.04	--
Iron concentration (mM)	40	40	40	--
Current density ( $\text{mA}/\text{cm}^2$ )	30	30	100	--

[0044] For each sample, a slurry was prepared with 6 mg of coal in 6 ml of IPA. The solution was then sonicated for 5 minutes. The slurry was then diluted 20 times using IPA and sonicated again. For each sample, 25 ml of the slurry was then deposited on a silicon wafer, and the solvent was allowed to evaporate.

[0045] Each wafer was moved to the furnace. The wafers were purged with argon at 1 SLPM for 20 minutes. The furnace was then ramped up to 800° C with argon flow for 30 minutes. The furnace was then maintained for 30 minutes at 800° C with a flow of a hydrogen-nitrogen

mixture at 0.8 SLPM. The system was then cooled down to room temperature with an argon flow of 1 SLPM.

**[0046]** As in Example 3 above, nanostructures grew from the wafer having raw coal, but the topography of the nanostructures did not follow the expected topography for graphene. The wafers with samples A-C had nanostructure growth exhibiting the topography of graphene. Sample C yielded the greatest amount of nanostructure growth, and sample A yielded the least amount of nanostructure growth. Accordingly, this example suggests that coal electrolyzed at a higher temperature (specifically, 104° C vs. 80° C) yields greater graphene synthesis, and that coal electrolyzed at a higher current density (specifically, 100 mA/cm<sup>2</sup> vs. 30 mA/cm<sup>2</sup>) yields greater graphene synthesis.

#### **Example 5**

**[0047]** In still another example, a slurry was prepared with 5.8 mg of electrolyzed coal in 6 ml isopropanol. The slurry was sonicated 5 minutes, and then diluted 20 times. A silicon wafer was cleaned with acetone and D-water, and then 100 µl of the slurry was deposited on the wafer. This same pre-treatment process was also used to prepare a second silicon wafer on which raw coal particles were deposited.

**[0048]** After drying, the wafers were positioned in the center of a quartz tube. The wafers were purged with 1000 SCCM of argon for 20 minutes. The furnace was then ramped up to 800° C with a 1000 SCCM argon flow for 30 minutes. The temperature of the system was then held at 800° C for additional 30 minutes with a 100 SCCM hydrogen-nitrogen flow (1:9). The system was then cooled down to room temperature under a 1000 SCCM argon flow.

**[0049]** An atomic force microscopy (AFM) image of the graphene sheets was taken under ambient conditions using a MFP-3D microscope (Asylum Research, Santa Barbara, CA) in AC mode. The AFM height image and 3D image showed the morphology of graphene sheets. The images and height measurement showed that the thickness was ~ 1 nm, indicating that the two-dimensional sheets were graphene monolayers. Large scale scanning showed that the lateral dimensions of the graphenes were up to a few hundred nanometers.

**[0050]** These Examples demonstrate that electrolyzed coal undergoes decomposition and graphenization at high temperatures. The hydrogen acts as both a reducing gas and a carrier gas, reacting with the graphitized coal at high temperatures to generate the hydrocarbons or other reactive intermediates. These hydrocarbons and reactive intermediates act as precursors for the synthesis of graphene on copper foil. This method is particularly effective and should significantly reduce the cost of forming high quality graphene for commercial use.

## **REFERENCES CITED IN THE DESCRIPTION**

Cited references

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**Patent documents cited in the description**

- [US7736475B \[0018\]](#)

**Non-patent literature cited in the description**

- Versatile Catalytic System for the Large-Scale and Controlled Synthesis of Single-Wall, Double-Wall, Multi-Wall, and Graphene Carbon Nanostructures CHEMISTRY OF MATERIALS, 2009, vol. 21, 225491-5498 [\[0003\]](#)

**Patentkrav**

1. Fremgangsmåde til at danne graphen, omfattende opvarmning af elektrolyseret kul til en temperatur, som er effektiv til at danne grafit under tilstedeværelse af en flydende strøm af reduktionsgas, idet temperaturen er mellem 490°C og 1100°C, hvor strømmen af reduktionsgas aflejrer graphen på en overflade, idet der dannes en graphen-film.  
5
2. Fremgangsmåde ifølge krav 1, hvor det elektrolyserede kul aflejres som et tyndt lag på et substrat, og reduktionsgassen føres over det tynde lag ved forhøjede temperaturer.
- 10 3. Fremgangsmåde ifølge krav 2, hvor det elektrolyserede kul omfatter et gelatinelignende materiale dannet på elektrolyserede kulpartikler.
4. Fremgangsmåde ifølge krav 2, hvor det elektrolyserede kul omfatter elektrolyserede kulpartikler belagt med en gelatinelignende film, som er dannet under elektrolyseprocessen.  
15
5. Fremgangsmåde ifølge krav 2, hvor det elektrolyserede kul omfatter elektrolyserede kulpartikler.
6. Fremgangsmåde ifølge krav 2, hvor reduktionsgassen omfatter hydrogen og en inert bæregas.  
20
7. Fremgangsmåde ifølge krav 6, hvor den inerte bæregas er nitrogen eller argon.
8. Fremgangsmåde ifølge krav 1, desuden omfattende at lade en strøm af inert gas flyde, før strømmen af reduktionsgas flyder.  
25
9. Fremgangsmåde ifølge krav 1, desuden omfattende hurtig afkøling af overfladen efter opvarmningstrinnet.
- 30 10. Fremgangsmåde ifølge krav 2, hvor substratet omfatter kobberfolie.
11. Fremgangsmåde ifølge krav 1, hvor overfladen omfatter kobberfolie.

**12.** Fremgangsmåde ifølge krav 1, hvor det elektrolyserede kul er blevet elektrolyseret ved en strømtæthed på over  $30 \text{ mA/cm}^2$  og under  $200 \text{ mA/cm}^2$ .

**13.** Fremgangsmåde ifølge krav 1, hvor det elektrolyserede kul er blevet elektrolyseret ved en  
5 temperatur lig med eller over  $80^\circ$  celsius.

**14.** Fremgangsmåde ifølge krav 4 eller 5, idet aflejringen af det elektrolyserede kul på  
substratet desuden omfatter dispergering af det elektrolyserede kul i et bærestof, brug af en  
gassprøjte til at afgive det elektrolyserede kul i bærestoffet på substratet, og fordampning af  
10 bærestoffet.

DRAWINGS

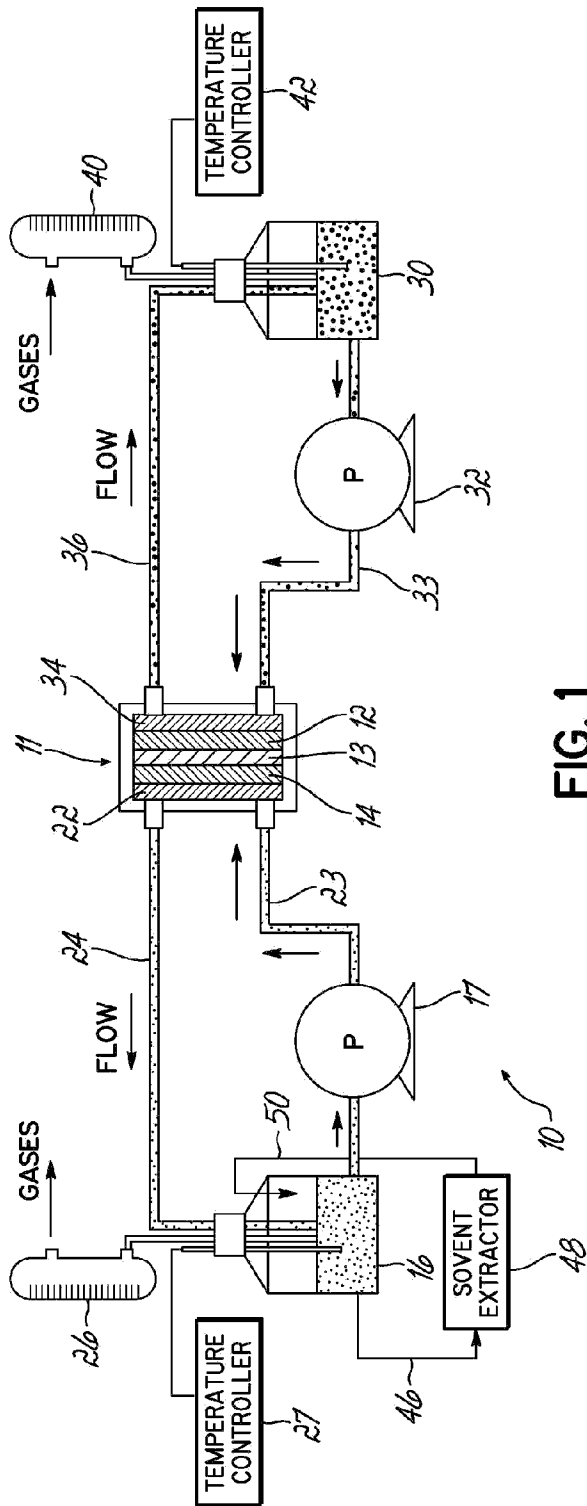


FIG. 1

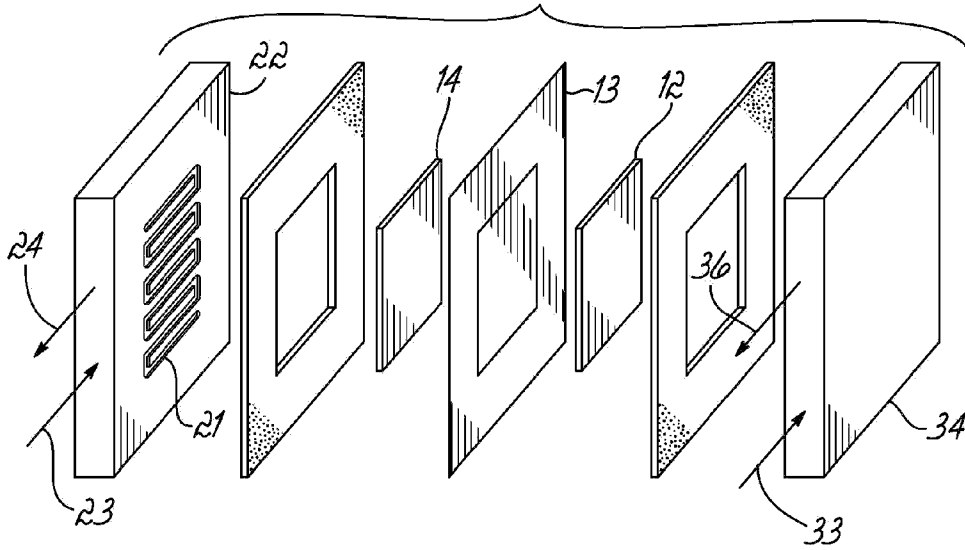


FIG. 2



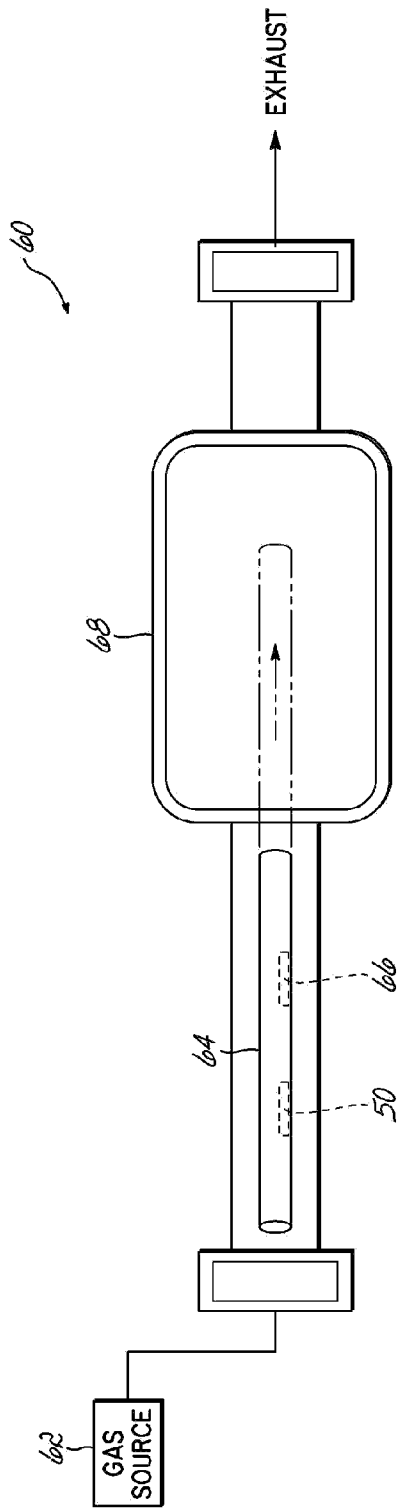


FIG. 3