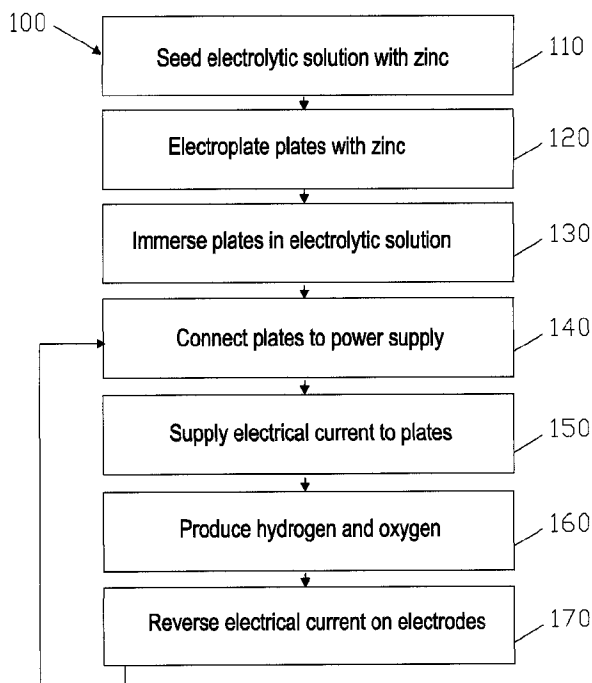




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 (72) Inventeur/Inventor:
 BURTCH, CHRISTOPHER J., CA
 (73) Propriétaire/Owner:
 BURTCH, CHRISTOPHER J., CA
 (74) Agent: NA

(54) Titre : PROCÉDE ET APPAREIL DE PRODUCTION D'HYDROGENE AYANT DES ELECTRODES REVERSIBLES
 (54) Title: METHOD AND APPARATUS FOR PRODUCING HYDROGEN HAVING REVERSIBLE ELECTRODES



(57) **Abrégé/Abstract:**

Provided is a method and an apparatus for producing hydrogen from salt water by electrolysis. The apparatus includes at least one pair of electrodes, wherein the pair of electrodes comprises a first zinc plated electrode and a second zinc plated electrode spaced apart from the first zinc plated electrode, a controller for supplying direct electrical current across the first and second electrodes such that the first and second electrodes are oppositely charged, and a tank containing an electrolytic solution including salt water and a mild acid. The first and second electrodes are immersed in the electrolytic solution.

Abstract

Provided is a method and an apparatus for producing hydrogen from salt water by electrolysis. The apparatus includes at least one pair of electrodes, wherein the pair of electrodes comprises a first zinc plated electrode and a second zinc plated electrode spaced apart from the first zinc plated electrode, a controller for supplying direct electrical current across the first and second electrodes such that the first and second electrodes are oppositely charged, and a tank containing an electrolytic solution including salt water and a mild acid. The first and second electrodes are immersed in the electrolytic solution.

**Title: METHOD AND APPARATUS FOR PRODUCING HYDROGEN HAVING
REVERSIBLE ELECTRODES**

Technical Field

[1] The embodiments disclosed herein relate to methods and apparatus for producing hydrogen, and, in particular to method and apparatus for producing hydrogen from salt water by electrolysis.

Background Art

[2] Electrolysis is a known method of producing hydrogen gas from water, in which two electrodes are placed in water and an electrical power supply is connected to the electrodes. The positively charged electrode is called the anode and the negatively charged electrode is the called the cathode. When performing fresh water electrolysis, water decomposes into hydrogen gas and oxygen gas. Hydrogen tends to form at the cathode and oxygen tends to form at the anode.

[3] One problem with the fresh water electrolysis is that fresh water has poor electrical conductivity, which prevents the flow of electrons from the electrical power supply, through the cathode, and to the anode. Without the flow of free electrons, the electrochemical reactions cannot occur, and water will not decompose into hydrogen and oxygen. For this reason, an electrolyte such as salt (e.g. sodium chloride) is added to the water so as to provide ions, which tends to increase the conductivity of water and improve the efficiency of electrolysis.

[4] While salt water electrolysis is more efficient than fresh water electrolysis, salt water electrolysis is still a relatively inefficient process for producing hydrogen. As a result, salt water electrolysis represents only a small fraction of the total hydrogen production worldwide.

[5] Increasing the efficiency of salt water electrolysis could substantially increase the amount of hydrogen produced worldwide. In particular, oceans are an abundant source of salt water and could be utilized to extract large amounts of hydrogen for use in commercial and industrial applications. However, this is not practical until salt water electrolysis becomes more efficient.

[6] United States patent number 8,282,812, to the same inventor, is directed to methods and apparatus for producing hydrogen from salt water by electrolysis, which uses a solid zinc anode and an aluminum cathode. The zinc electrode deteriorates and is replaced, while the zinc oxide, being produced, is continually removed. Zinc is consumed during the process and zinc oxide is produced as well as hydrogen. This necessitates the replacement of the zinc plates and the removal of the zinc oxide.

[7] Accordingly, there is a need for new or improved apparatus and methods for producing hydrogen from salt water by electrolysis.

Summary of Invention

[8] According to some embodiments, there is an apparatus for producing hydrogen from salt water by electrolysis. The apparatus comprises at least one pair of electrodes, wherein the pair of electrodes comprises a first zinc plated electrode and a second zinc plated electrode spaced apart from the first zinc plated electrode, a controller for supplying direct electrical current across the first and second electrodes such that the first and second electrodes are oppositely charged, and a tank containing an electrolytic solution including salt water and a mild acid. The first and second electrodes are immersed in the electrolytic solution.

[9] The controller may include a switch for reversing the direction of electrical current flow supplied across the first and second electrodes such that the first electrode switches from being an anode to being a cathode and the second electrode switches from being a cathode to being an anode.

[10] The first electrode and the second electrode may each comprise an aluminum plate.

[11] The second electrode may be thinly plated with zinc such that the aluminum plate does not oxidize in the electrolytic solution.

[12] The thin plating of zinc may be about 1/16 of the thickness of the aluminum plate of the second electrode.

[13] The first electrode may be thickly plated with zinc such that the thickness of the zinc plating is dependent on a desired run time.

- [14] The plating time may be about three times the desired run time.
- [15] The thick plating of the zinc may be about $\frac{1}{4}$ of the thickness of the aluminum plate of the first electrode.
- [16] The mild acid may be vinegar.
- [17] The electrolytic solution may be seeded with zinc.
- [18] The first and second electrodes may include any one of aluminum alloy, copper, or steel.
- [19] The apparatus may further comprise a collection pipe fitted into the lid in order to collect the gases produced at the anode and cathode of the apparatus and direct them to their proper locations.
- [20] The collection pipe may include a graphene filter to separate the hydrogen gas from the oxygen gas and water vapor.
- [21] The collection pipe may include vents for the oxygen and water vapor to be released.
- [22] The controller may include a regulator for regulating the amount of electrical voltage and current supplied across the first and second electrodes.
- [23] The controller may include a sensor for monitoring the electrical current. The regulator may regulate the electrical current based on the sensed electrical current.
- [24] The controller may reverse the direction of the electrical current flow after the desired run time.
- [25] Each pair of electrodes may comprise one thickly zinc coated anode and one thinly zinc coated cathode.
- [26] According to some embodiments, there is a method of producing hydrogen from salt water by electrolysis. The method comprises immersing at least one pair of electrodes in an electrolytic solution including salt water and a mild acid, wherein the pair of electrodes comprises a first zinc plated electrode and a second zinc plated electrode spaced apart from the first zinc plated electrode, supplying direct electrical current across the first and second electrodes such that the first and second electrodes

are oppositely charged, collecting the hydrogen gas produced at the second electrode, collecting the oxygen gas produced at the first electrode, and reversing the direction of the electrical current supplied across the first and second electrodes.

[27] The method may further comprise thickly plating the first electrode with zinc such that the thickness of the zinc plating is dependent on a desired run time.

[28] The method may further comprise thinly plated the second electrode with zinc such the thin plating of zinc is about 1/16 of the thickness of the second electrode.

[29] The method may further comprise seeding the electrolytic solution with zinc.

[30] The method may further comprise regulating the amount of electrical voltage and current supplied across the first and second electrodes.

[31] The method may further comprise sensing the electrical current, and wherein regulating the electrical current based on the sensed electrical current.

[32] The method may further comprise reversing the direction of the electrical current supplied across the first and second electrodes after the desired run time, wherein the run time is based on the amount of zinc plated on the first electrode.

[33] Other aspects and features will become apparent, to those ordinarily skilled in the art, upon review of the following description of some exemplary embodiments.

Brief Description of the Drawings

[34] The drawings included herewith are for illustrating various examples of articles, methods, and apparatuses of the present specification. In the drawings:

[35] Figure 1 is a section view of an apparatus for producing hydrogen, in accordance with an embodiment;

[36] Figure 2 is a section view of the apparatus of Figure 1 in a second position;

[37] Figure 3 is a side view of a cathode and an anode of the apparatus of Figure 1;

[38] Figure 4 is a flow chart of a method for producing hydrogen, in accordance with an embodiment.

Detailed Description

[39] Various apparatuses or processes will be described below to provide an example of each claimed embodiment. No embodiment described below limits any claimed embodiment and any claimed embodiment may cover processes or apparatuses that differ from those described below. The claimed embodiments are not limited to apparatuses or processes having all of the features of any one apparatus or process described below or to features common to multiple or all of the apparatuses described below.

[40] Referring to Figure 1, illustrated therein is an apparatus 3 for producing hydrogen and oxygen from salt water, in accordance with an embodiment. The apparatus 3 includes at least one pair of electrodes. Each pair of electrodes includes a first electrode 1 and a second electrode 2 that are oppositely charged and arranged in a stack. The first and second electrodes 1, 2 are separated apart from each other and immersed in a tank 17 containing an electrolytic solution 26. As shown, the apparatus 3 includes four pairs of electrodes 1, 2. The first electrode 1 and the second electrode 2 are connected to a power supply 29 and controller 36 so as to be oppositely charged to produce hydrogen and oxygen gasses.

[41] As shown, the first electrode 1 may be an anode that is a positively charged electrode that attracts electrons or negatively charged atoms. The second electrode 2 may be a cathode that is a negatively charged electrode that attracts protons or positively charged atoms. The apparatus 3 produces hydrogen and oxygen from salt water in the electrolytic solution 26 by electrolysis and the oxidization of zinc.

[42] Each of the first electrodes 1 are switched between operating as an anode (Figures 1 and 3) and a cathode (Figures 2 and 3). Each of the second electrodes 2

are correspondingly switched between operating as a cathode (Figure 1), and an anode (Figure 2).

[43] The electrodes 1, 2 are metal plates plated with zinc. In an embodiment, the first electrode 1 is made of an aluminum plate and is initially heavily coated in zinc while acting as the anode. The second electrode 2 is also made of an aluminum plate and is initially thinly coated in zinc while acting as the cathode.

[44] The electrolytic solution 26 includes a mixture of salt water and a mild acid, such as vinegar. The electrolytic solution 26 allows the oxygen to be released at the anode and hydrogen to be released at the cathode. As the zinc oxidizes in the electrolytic solution 26, the mild acid breaks down the zinc oxide that is produced. The broken down zinc migrates to the cathode where the zinc is electro-plated to the surface of the cathode. The zinc coating on the anode gradually decreases as the zinc oxidizes and the zinc coating on the cathode gradually increases as the zinc is electro-plated onto the cathode.

[45] The mild acid may be vinegar or sulphuric acid. Vinegar may be particularly desirable as it may be inexpensive, effective, and readily available. The mild acid facilitates the separation of oxide produced back into zinc and oxygen. The mild acid is mixed with a much larger portion (e.g., between 6:1 and 10:1) of salt water to form the electrolytic solution 26. The mild acid dissolves the zinc oxide, as the zinc oxide is produced, and releases oxygen onto the anode while the zinc is plated to the cathode. Without the mild acid and the separation of the zinc oxide there would be the creation of zinc oxide and the consumption of zinc.

[46] In an embodiment, the electrolytic solution 26 is seeded with zinc. The zinc is seeded into the mild acid solution. In an embodiment, the electrolytic solution 26 is seeded with zinc by placing a sheet of zinc into the vinegar before the mild acid is added to the salt water. After the zinc has dispersed into the mild acid over a period of time (e.g., a few hours), the sheet of zinc is removed from the mild acid and the seeded mild acid solution is mixed with the salt water. The zinc seeding allows for faster transfer of plated zinc through the electrolytic solution 26.

[47] In an embodiment, the electrolytic solution 26 is prepared and seeded once at the beginning of the process. The electrolytic solution 26 may contain about eight (8) parts of salt water to one (1) part of vinegar with zinc dissolved (seeded) into the mixture.

[48] The zinc may be seeded into the electrolytic solution by dissolving zinc into undiluted vinegar for a period of time (e.g., 2 hours). The vinegar is then stirred into the salt water and a sheet of zinc is placed into the mixture for a period of time (e.g., 24 hours) in order to further seed the electrolyte. During production the water and vinegar are slowly consumed and are replenished in order to top up the level of the electrolyte and no further zinc is added.

[49] The anode releases oxygen and zinc while the cathode releases hydrogen and attracts zinc. The acid breaks apart the zinc oxide while the oxidization enables less current use than regular electrolysis, in order to separate the water molecules into hydrogen and oxygen. Oxidization of the zinc allows for the production of hydrogen from the salt water which may use less electrical current than prior art methods of electrolysis.

[50] The anode may begin the process as thickly plated with zinc and the cathode begins as thinly plated with zinc.

[51] Initially, the cathode may be thinly plated with zinc, with just enough of a coating so that the aluminum plate does not oxidize in the electrolytic solution 26. The anode is plated with zinc to a thickness dependent on a plating time and a desired run time. In some cases, the anode and cathode may be of equal zinc plating thickness, however, once the electrolysis begins, the cathode will acquire zinc from the anode.

[52] As the electrolysis occurs, the zinc passes from the anode to the cathode. After a period of time, the charges on the electrodes 1, 2 are reversed (see Figure 2), so that the first electrode 1 becomes a cathode and the second electrode 2 becomes an anode. When the current flow is reversed across the electrodes 1, 2, the zinc migrates in the opposite direction.

[53] In an embodiment, the electrodes 1, 2 are made of 5000 series aluminum plates that are resistant to salt water deterioration. The aluminum plates are electroplated with a layer of zinc, which provides corrosion resistance and electrical conductivity. In an embodiment, the aluminum plates are 5052 grade aluminum, which may be particularly resistant to corrosion in salt water.

[54] Aluminum may be particularly advantageous as zinc electroplates to aluminum and aluminum is a good conductor of electricity. In other embodiments, the electrodes 1, 2 may be made from another material such as zinc plated metal such as copper, platinum, steel, stainless steel, a steel alloy, or another conductive material with corrosion resistance, electrical conductivity and the ability to be zinc plated.

[55] The electrodes 1, 2 may be four-inch square plates having a thickness of approximately 1/32 of an inch. In other embodiments, the electrodes 1, 2 may have different sizes and shapes.

[56] In an embodiment, the thin plating of zinc is about 1/16 of the thickness of the aluminum plate in the cathode. The thick plating of the zinc is about 1/4 of the thickness of the aluminum plate of the anode.

[57] For example, the aluminum plates of the electrodes 1, 2 may be 1/16" in thickness. When aluminum plate of the cathode is 1/16" thick, the thin coating of zinc is about 1/256" thick. When aluminum plate of the anode is 1/16" thick, the thick coating of the zinc is about 1/64" thick up to a thickness of 1/32". The thickness of the zinc coating is dependent on the length of the desired run time. The plating time may be about three times the desired run time. For example, if the desired run time is 1 hour, before switching current, then the anode is plated for 3 hours. Further, the anode should not be plated so thick as to flake off of the plate or to close the spacing to the cathode. The anode and cathode should not touch.

[58] In a particular embodiment, shown in Figures 1 and 2, the apparatus 3 includes four pairs of electrodes 1, 2, four anodes and four cathodes. The anodes and cathodes are arranged in alternating sequence such that each cathode is spaced apart from an adjacent anode (e.g. one cathode followed by one anode). In other

embodiments, there may be a different number of electrodes 1, 2, for example, at least one anode and at least one cathode.

[59] In an embodiment, the apparatus 3 includes a tapper 43 (shown schematically) for tapping the electrodes 1, 2 to disperse gas bubbles, at the end of each cycle before the electrical current is switched.

[60] During a run phase, the current is applied across the electrodes 1, 2, and hydrogen forms on the cathode while the zinc coating on the cathode increases in thickness. Hydrogen is released at the cathode and oxygen is released at the anode. Once the zinc coating on the anode becomes thin, the electrical current across the anode and cathode is reversed so that the now thickly coated cathode becomes an anode while the thinly coated anode becomes the cathode. The current is applied across the anode and cathode so that the initial arrangement is reversed and one plate is thinly coated and the other plate is thickly coated.

[61] The apparatus 3 includes the controller 36 for controlling the electrical current and voltage being supplied by the power supply 29. The controller 36 includes a regulator 28 for regulating the amount of electrical voltage and current supplied as the amount of current changes continuously while the anodes and cathodes sacrifice and accumulate zinc, respectively. The current varies based on the thicknesses of the zinc and the electrolytic solution 26.

[62] The controller 36 may also include a sensor 38 for sensing the electrical current across the electrodes 1, 2. The sensor 38 detects the amount of current flow as the zinc travels in the electrolytic solution 26. The regulator 28 regulates the electrical current based on the sensed electrical voltage and current to keep the electrical current in a preferred range. For example, where there is a voltage drop, the sensor 38 detects the voltage drop and the regulator 28 increases the voltage. The preferred voltage range is based on the size and the scale of the apparatus 3.

[63] The apparatus 3 may also include a pump 40 for circulating the electrolytic solution. The pump 40 may also circulate added salt water and mild acid as the salt water is consumed. The pump 40 may also prevent bubbles of oxygen and hydrogen from forming on the electrodes 1, 2.

[64] The controller 36 may include a switch 37 for periodically reversing the direction of electrical current flow. In certain cases, the switch 37 automatically reverses the direction of the current flow based on time and/or thickness of the electrodes 1, 2. The controller 36 may include a timer 39 for determining a run time set to trigger the switch 37 to reverse the direction of the electrical current flow. The run time may be based on the amount of zinc plated on the anode.

[65] Referring now to Figure 2, illustrated therein is the apparatus 3' after the electrical current has been reversed. The first electrode 1' is a cathode and the second electrode 2' is an anode. The first electrode 1' releases hydrogen and is electro-plated with zinc from the second electrode 2'. The second electrode 2' loses zinc and releases oxygen onto the surface of the second electrode 2'. The electrical current can be reversed continuously, consuming only the salt water and vinegar, with the zinc being electro-plated from the anode to the cathode. The heavier zinc deposit is located on the anode. The cathode is connected to the negative terminal of a direct current power supply 29, through the regulator 28, which indirectly connects the other cathodes to the power supply 29.

[66] Referring again to Figure 1, the plurality of second electrodes 2 are electrically interconnected to each other in parallel by second electrode couplers 8. For example, the second electrode couplers 8 may be push on clips that can be used to form a continuous link between the second electrodes 2. The push on clips include an appropriate gauge wire 9 having two female nylon clips at the ends of the wire 9 that slide over two respective second electrodes 2. The wire 9 is crimped inside the clips and is in electrical contact with each second electrode 2 so as to electrically interconnect the two second electrodes 2. In other embodiments, the second electrode couplers 8 may be connected to the second electrodes 2 in different ways, for example, by soldering the wires to the second electrodes 2.

[67] The apparatus 3 includes a second electrode end connector 31 for electrically connecting the second electrodes 2 to the power supply 29. The second electrode end connector 31 includes a current supply wire 18, having an appropriate gauge, with a female nylon clip on one end of the wire, and a male disconnect at the

other end of the wire. The female nylon clip can slide onto the second electrode 2 and the male disconnect can be connected to the power supply 29. In other embodiments, the second electrodes 2 may be connected to the power supply 29 in different ways, for example, using other types of end connectors.

[68] The plurality of first electrodes 1 are electrically interconnected to each other in parallel by first electrode couplers 7. For example, the first electrode couplers 7 may be push on couplers that can be used to form a continuous link between the first electrodes 1. The first electrode couplers 7 are generally similar to the second electrode couplers 8 and include an appropriate gauge wire with two female nylon clips. In other embodiments, the first electrode couplers 7 may be connected to the first electrodes 1 in different ways, for example, by soldering the wires to the first electrodes 1.

[69] One of the first electrodes 1 is connected to a positive terminal of the direct current power supply 29, through the regulator 28, which indirectly connects the remaining first electrodes 1 to the direct current power supply 29. The apparatus 3 includes a first electrode end connector 32 for electrically connecting the first electrodes 1 to the power supply 29. For example, the first electrode end connector 32 may be generally similar to the second electrode end connector 31 and may include a current supply wire 19, having an appropriate gauge, with a female nylon clip and a male disconnect. In other embodiments, the first electrodes 1 may be connected to the power supply 29 in different ways, for example, using other types of end connectors.

[70] The apparatus 3 also includes non-conductive spacers 42 coupled to the electrodes 1, 2 for spacing apart the electrodes 1, 2 in an electrically insulated manner. For example, the non-conductive spacers 42 may include insulating bolts 4, insulating nuts 5 and insulating washers 6 for fastening the electrodes 1, 2 together within the stack. The non-conductive spacers 42 electrically insulate the electrodes 1, 2 from each other during electrolysis, while also positioning the electrodes 1, 2 within the stack.

[71] The non-conductive spacers 42 may removably couple the electrodes 1, 2, for example, such that the electrodes 1, 2 can be replaced. For example, the insulating nuts 5 may be unthreaded from the insulating bolts 4 such that the electrodes 1, 2, and the insulating washers 6 can be removed and then replaced.

[72] The apparatus 3 includes a collection system positioned above the electrodes 1, 2 for collecting the produced oxygen and hydrogen gases. During electrolysis, hydrogen bubbles form around the cathodes while oxygen bubbles form around the anodes. The oxygen and hydrogen thus produced rises in the electrolyte 26 to the upper surface and they are then separated and collected.

[73] The gases exit the apparatus 3 by an opening 51 in the lid 48 where a pipe is fitted into the lid allowing for an upward flow of the gases.

[74] The apparatus lid 48 is firmly attached to the lower housing 17 using a gasket 49 between the two parts and a latch 50 to lock the lid tightly in place.

[75] The gases rise to the graphene filter 46. A recently discovered property of graphene is that it will allow hydrogen to pass through it while blocking all other gases and water vapor.

[76] The hydrogen proceeds via a pipe 45 to a compressor and storage tank not illustrated herein.

[77] The other gases and water vapor produced that cannot pass through the graphene filter 46 are vented through holes 47 in the feeder pipe.

[78] The gas that passes through the graphene filter is thus 100% hydrogen with no other gases or contaminants present in the final product.

[79] Referring now to Figure 3, illustrated therein are two electrodes, the top illustration being the anode 1 from Figure 1 and the bottom illustration being the cathode from Figure 1. The first electrode 1 starts off connected to the positive terminal, and loses its zinc coating while producing oxygen. The second electrode 2 is connected to the negative terminal of a direct current power supply and gains a zinc coating while producing hydrogen. Eventually the current is reversed as the zinc deposit on the anode 1 becomes too thinly deposited.

[80] The chemical reactions that produce both the hydrogen and oxygen may occur on the surface of the electrodes 1, 2. It may be desirable to increase the surface area of the electrodes 1, 2 so as to increase the amount of gases that are produced.

Accordingly, the electrodes 1, 2 may include a multiplicity of apertures 27 for increasing the surface area of the electrodes 1, 2.

[81] In some embodiments, the apertures 27 may have a diameter of approximately 1/16 of an inch and a density of approximately 64 apertures per square inch. The apertures 27 may be distributed in a pattern so as to provide perforated electrodes 1, 2. While the apertures 27 of the illustrated embodiment are circular, the apertures 27 may be other shapes such as square or oval. The zinc plating on the electrodes 1, 2 is such that the apertures 27 continue to be present.

[82] The electrodes 1, 2 include four bolt holes 21 for receiving four of the insulating bolts 4. The insulating bolts 4 may be made of an insulating material such as plastic. The insulating bolts 4 cooperate with the insulating nuts 5 and insulating washers 6 to couple the electrodes 1, 2 together in an electrode stack while electrically insulating oppositely charged electrodes 1, 2. In other embodiments, there may be a different number of bolt holes 21 depending on the desired number of insulating bolts 4.

[83] The second electrode couplers 8 electrically interconnect the set of second electrodes 2 and the first electrode couplers 7 electrically interconnect the set of first electrodes 1. To facilitate the use of couplers 7, 8, each electrode 1, 2 may have tabs 20 cut into the side of the plate for receiving the couplers 7, 8.

[84] The first electrode 1 may have two tabs 20 on the one side for receiving two different first electrode couplers 7. In particular, one tab 20 may receive the female nylon clip of the first electrode coupler 7 that is coupled to another similarly charged electrode 1, and the other tab 20 may receive the female nylon clip of a different first electrode coupler 7 that is coupled to a further similarly charged electrode 1.

[85] Similarly, the second electrode 2 has two tabs 20 on the opposite side to that of the first electrodes 1. Positioning the tabs 20 of the second electrode 2 on the opposite side as the tabs 20 of the first electrode 1 may reduce interference between the second electrode couplers 8 and the first electrode couplers 7.

[86] In Figure 1 the set of second electrodes 2 are connected in parallel to the terminal of the direct current power supply 29 through the regulator 28. In particular, the

second electrode 2 closest to the nut end (e.g. the left side of Figure 1) is connected to the power supply 29 using the second electrode end connector 31. The second electrode end connector 31 includes the current supply wire 18 having a female nylon clip on one end and a male disconnect on the other end. The female nylon clip slides over one of the tabs 20 on the second electrode 2. The male disconnect can be connected to the direct current power supply 29 through a corresponding female connector 11 on a wire 13 connected to the terminal of the direct current power supply 29.

[87] Similarly, the set of first electrodes 1 are connected in parallel to the opposite terminal of the direct current power supply 29. In particular, the first electrode 1 closest to the nut end (e.g. the left side of Figure 1) is connected to the power supply 29 using the first electrode end connector 32. The first electrode end connector 32 includes a current supply wire 19 having a female nylon clip on one end and a male disconnect on the other end. The female nylon clip slides over one of the tabs 20 on the first electrode 1. The male disconnect can be connected to the direct current power supply 29 through a corresponding opposite female connector 10 on a wire 12 connected to the opposite terminal of the direct current power supply 29.

[88] Tests were conducted using the electrode stack illustrated in Figures 1 and 2, including a set of four zinc coated perforated aluminum first electrodes 1 and a set of four oppositely charged zinc coated perforated aluminum electrodes 2. The electrodes 1, 2 being four inches square (10.16 cm) and have a thickness of approximately 1/32 of an inch (0.792 mm).

[89] It has been determined that supplying approximately one volt of electrical potential across the electrodes 1, 2 in the stack at approximately one amp (e.g. approximately one watt of power) may provide an efficient means for producing hydrogen. For example, the apparatus 3 produced approximately 400 milliliters of hydrogen gas every hour using 1 watt of electricity.

[90] Based on this data, it may be desirable to configure the power supply 29 and regulator 28 to provide between approximately 0.8 and 1.2 volts of electrical

potential and between approximately 0.8 and 1.2 amperes, in order to maximize the hydrogen gas produced per watt of power supplied.

[91] In some embodiments, there may be a plurality of electrode stacks supplied with approximately 0.8 and 1.2 volts and between approximately 0.8 and 1.2 amperes. Providing a plurality of electrode stacks may increase the amount of hydrogen produced by this process.

[92] In some embodiments, there may be a different number of electrodes 1, 2 in the electrode stack and the amount of power applied to each electrode stack may be different. Accordingly, in some embodiments, the power supply 29 may be configured to supply between approximately 0.8 and 1.2 volts and an amperage based on the number of electrodes 1, 2 in the electrode stack. For example, the amperage may be between approximately 0.1 and 0.4 amperes for each second electrode 2 and each first electrode 1.

[93] While the illustrated embodiment of Figures 1 to 3 includes electrodes 1, 2 having a particular size, shape and surface area (e.g. approximately sixteen square-inches, not accounting for the plate thickness, apertures 27 and bolt holes 21), other embodiments may include electrodes 1, 2 having a different size, shape and surface area.

[94] In some embodiments, each plate may have a different size, shape, or surface area, and the amount of power applied to each electrode 1, 2 may be different. Accordingly, in some embodiments, the power supply 29 may be configured to supply between approximately 0.8 and 1.2 volts and a current density based on the number of the plates in the stack and the size, shape and surface area of each plate. For example, the current density may be between approximately 0.003 and 0.03 amperes per square-inch for each electrode 1, 2.

[95] Referring now to Figure 4, illustrated therein is a method 100 of salt water electrolysis, in accordance with an embodiment. The method 100 is for producing hydrogen and oxygen from salt water electrolysis in an electrolyte of salt water and vinegar, and may use the apparatus 3 as described with reference to Figures 1 to 3.

[96] At 110, the electrolytic solution is seeded with zinc.

[97] At 120, a heavy zinc coating is electroplated onto half of the electrode plates and a thin zinc coating is applied onto the other half of the electrode plates used in the electrode stack to create cathode plates (e.g., second electrodes 2 of Figure 1) and anode plates (e.g., first electrodes 1 of Figure 1). During a set up phase, one or more aluminum plates are thickly plated with zinc to create the anode, and one or more aluminum plates are thinly plated with zinc to create the cathode. The anode is thickly plated with zinc such that the thickness of the zinc plating is dependent on a desired run time. The cathode is thinly plated with zinc such the thin plating of zinc is about 1/16 of the thickness of the cathode plate.

[98] The zinc plating prevents the aluminum from oxidizing. The anode is electroplated with zinc for approximately three (3) times as long as the desired run time. For example, if the desired run time is 1 hours, the anode is plated for 3 hours.

[99] At 130, a stack of zinc coated plates is immersed into an electrolytic solution including salt water and mild acid such as vinegar. The electrode stack includes at least one zinc plated anode and at least one zinc plated cathode spaced apart from the anode. The anode plate is the heavily zinc coated plate and the cathode plate is the lightly zinc coated plate.

[100] At 140, the anode plate is connected to the positive terminal of a direct current power supply and the cathode plate is connected to the negative terminal of the same power supply.

[101] At 150, direct electrical current is supplied across the zinc plated anode and the zinc plated cathode. The cathode and anode may be provided between approximately 0.8 and 1.2 volts and between approximately 0.8 and 1.2 amperes when using the apparatus 3 described above.

[102] At 160, hydrogen gas is produced at the cathode and oxygen gas is produced at the anode and they exit the apparatus by a pipe connected to the lid. The gases then continue to rise, are separated, and the hydrogen is collected and stored while the oxygen is vented.

[103] The zinc coating on the aluminum anode plates gradually becomes thinner while the zinc coating on the perforated aluminum cathode plates becomes thicker. The water and the vinegar are consumed during this process, and are replenished in the electrolytic solution in order to keep the concentration and depth of the electrolytic solution at optimum operational levels.

[104] At 170, the direction of the electrical current is reversed once the zinc coating on the anode plate has become quite thin. The power supply is turned off and the anodes are switched by connecting the anodes to the negative terminal of the direct current power supply and the former cathodes are connected to the positive terminal making the plate function as the anode plate of the apparatus 3.

[105] After the zinc coating on the anode plate becomes thin then the cycle is resumed by disconnecting the power and proceeding to 140 in order to continuously produce hydrogen and oxygen while accommodating the migration of the zinc from the anode plate to the cathode plate.

[106] In some embodiments the method 100 may be used with an apparatus having plates with different sizes and shapes than those described and illustrated herein.

[107] While the above description provides examples of one or more apparatus, methods, or systems, it will be appreciated that other apparatus, methods, or systems may be within the scope of the claims as interpreted by one of skill in the art.

Claims

- Claim 1. An apparatus for producing hydrogen from salt water by electrolysis, the apparatus comprising:
at least one pair of electrodes, wherein the pair of electrodes comprises a first zinc plated electrode and a second zinc plated electrode spaced apart from the first zinc plated electrode;
a controller for supplying direct electrical current across the first and second electrodes such that the first and second electrodes are oppositely charged; and
a tank containing an electrolytic solution including salt water and a weak acid with a PH between 2 and 3;
wherein the first and second electrodes are immersed in the electrolytic solution.
- Claim 2. The apparatus of claim 1, the controller includes a switch for reversing the direction of electrical current flow supplied across the first and second electrodes such that the first electrode switches from being an anode to being a cathode and the second electrode switches from being a cathode to being an anode.
- Claim 3. The apparatus of claim 1, wherein the first electrode and the second electrode each comprise an aluminum plate.
- Claim 4. The apparatus of claim 3, wherein the second electrode is plated with zinc about 1/16 of the thickness of the aluminum plate so that the aluminum plate does not oxidize in the electrolytic solution.
- Claim 5. The apparatus of claim 4, wherein the plating of the zinc completely covers the aluminum electrode so that no oxidization of the aluminum will take place in the electrolyte.

- Claim 6. The apparatus of claim 4, wherein the first electrode is thickly plated with zinc such that the thickness of the zinc plating is dependent on a desired run time.
- Claim 7. The apparatus of claim 6 having a zinc coating applied for a longer duration in its application than the length of the desired run time for the apparatus in claim 1.
- Claim 8. The apparatus of claim 6, wherein the thick plating of the zinc is about $\frac{1}{4}$ of the thickness of the aluminum plate of the first electrode.
- Claim 9. The apparatus of claim 1, wherein the mild acid is vinegar.
- Claim 10. The apparatus of claim 1, wherein the electrolytic solution is seeded with zinc.
- Claim 11. The apparatus of claim 1, wherein the first and second electrodes include any one of aluminum alloy, copper, or steel.
- Claim 12. The apparatus of claim 1 further comprising a collection system to collect, filter and direct the gases produced at the electrodes.
- Claim 13. The apparatus of claim 12 wherein the collection system includes a collector pipe for the gases produced and a graphene or palladium filter to separate the hydrogen from any other gases or water vapor produced.
- Claim 14. The apparatus of claim 13, where the collection system is configured to release oxygen and water vapour through holes in the collection pipe and to collect hydrogen in a tank on the filtrate side of the filter.

- Claim 15. The apparatus of claim 1, wherein the controller includes a regulator for regulating the amount of electrical voltage and current supplied across the first and second electrodes.
- Claim 16. The apparatus of claim 15, wherein the controller includes a sensor to determine the voltage and amperage that is being supplied to the electrodes, so that they can be adjusted to the correct levels throughout the process.
- Claim 17. The apparatus of claim 16, wherein the controller can reverse the direction of the current flow, once the desired running time has been achieved, so that the anode then becomes the cathode and the cathode then becomes the anode.
- Claim 18. The apparatus of claim 1, wherein each pair of electrodes comprises one thickly zinc coated anode and one thinly zinc coated cathode.
- Claim 19. A method of producing hydrogen from salt water by electrolysis, the method comprising:
immersing at least one pair of electrodes in an electrolytic solution including salt water and a mild acid, wherein the pair of electrodes comprises a first zinc plated electrode and a second zinc plated electrode spaced apart from the first zinc plated electrode;
supplying direct electrical current across the first and second electrodes such that the first and second electrodes are oppositely charged;
collecting the hydrogen gas produced at the second electrode;
collecting the oxygen gas produced at the first electrode; and
reversing the direction of the electrical current supplied across the first and second electrodes.

- Claim 20. The method of claim 19 further comprising thickly plating the first electrode with zinc such that the thickness of the zinc plating is dependent on a desired run time.
- Claim 21. The method of claim 19 further comprising thinly plating the second electrode with zinc such that the thin plating of zinc is about 1/16 of the thickness of the second electrode.
- Claim 22. The method of claim 19 further comprising seeding the electrolytic solution with zinc.
- Claim 23. The method of claim 22 further comprising regulating the amount of electrical voltage and current supplied across the first and second electrodes.
- Claim 24. The method of claim 23 further comprising sensing the electrical current, and wherein regulating the electrical current is based on the sensed electrical current.
- Claim 25. The method of claim 23 further comprising reversing the direction of the electrical current supplied across the first and second electrodes after the desired run time, wherein the run time is based on the amount of zinc plated on the first electrode.

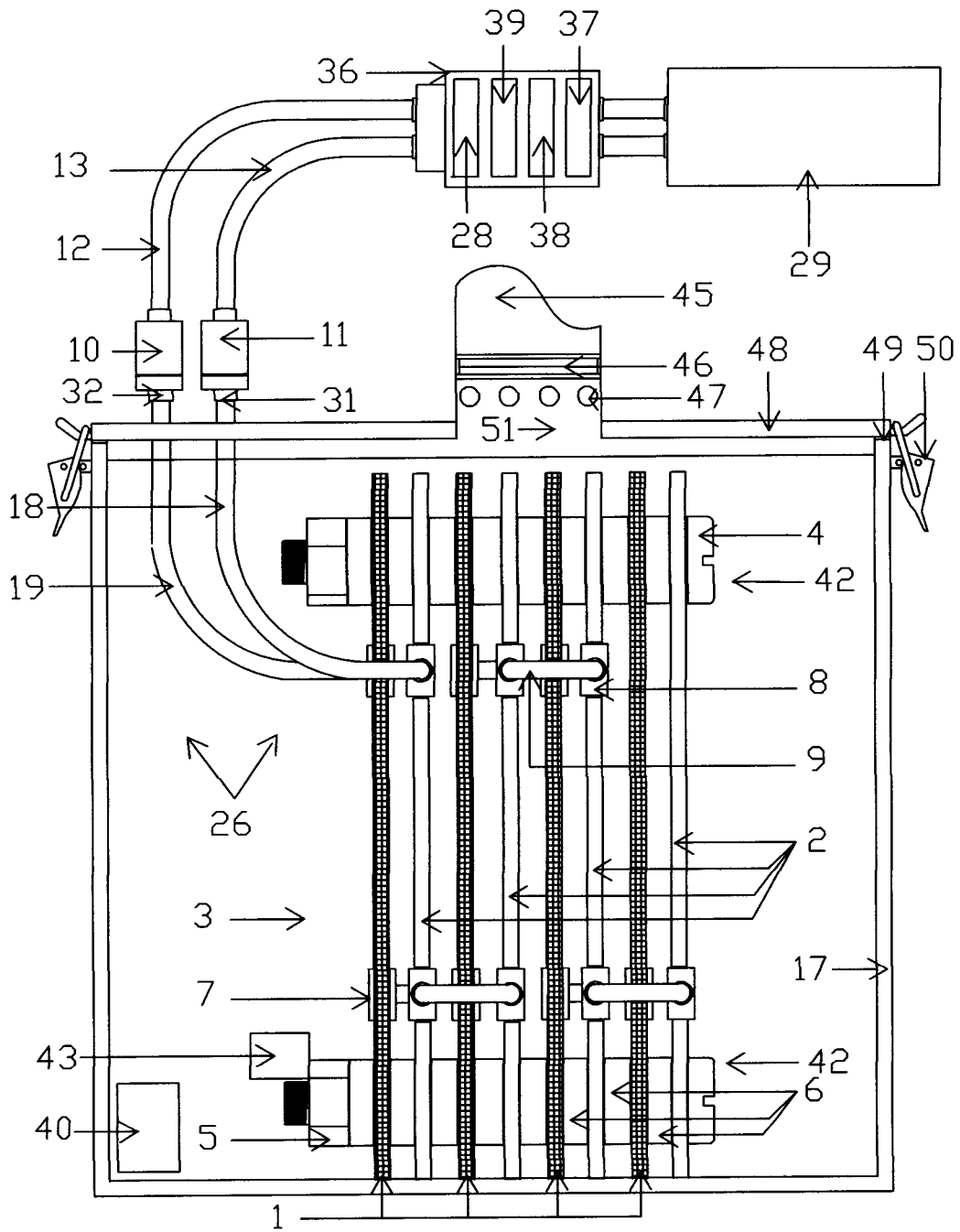


Figure 1

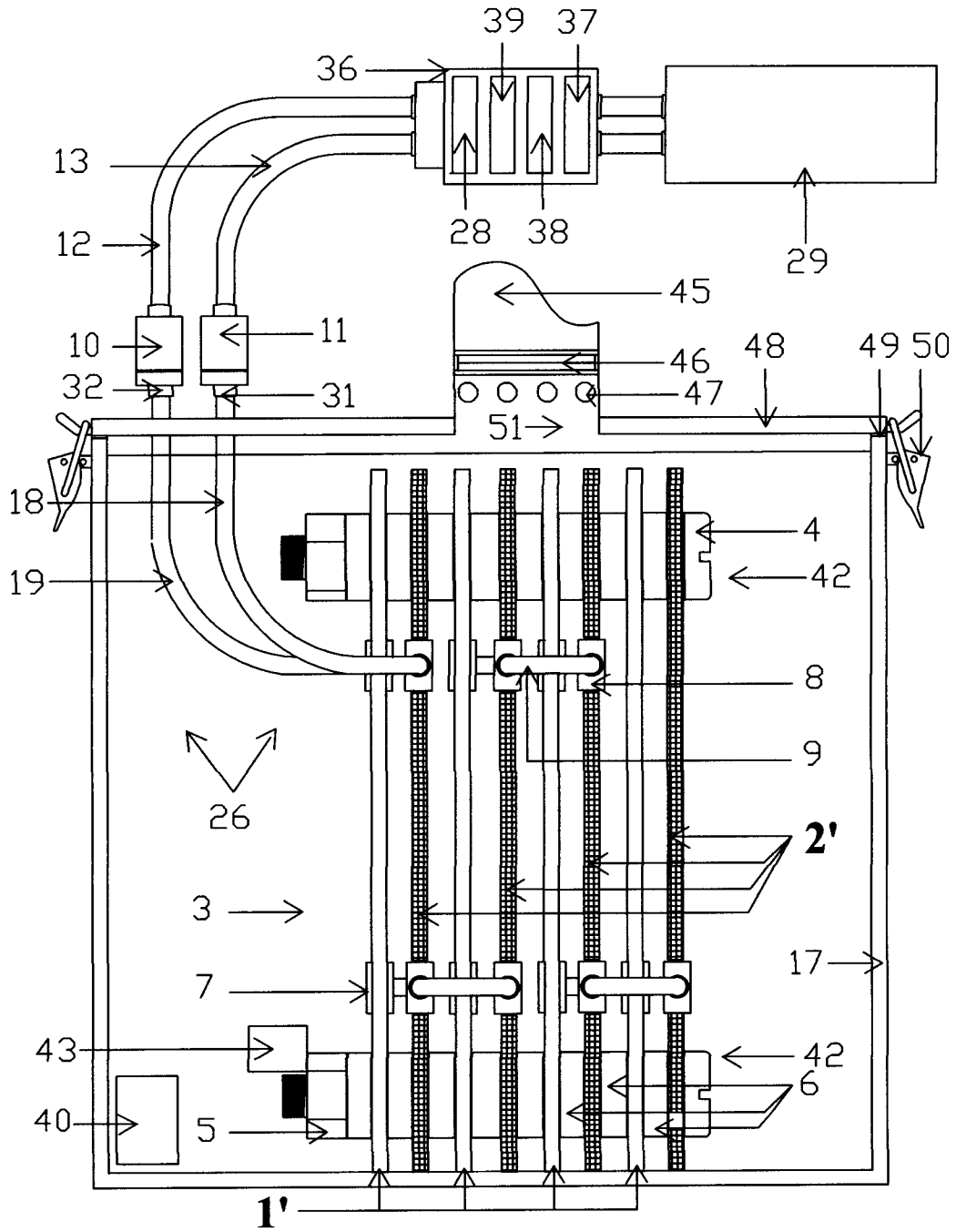
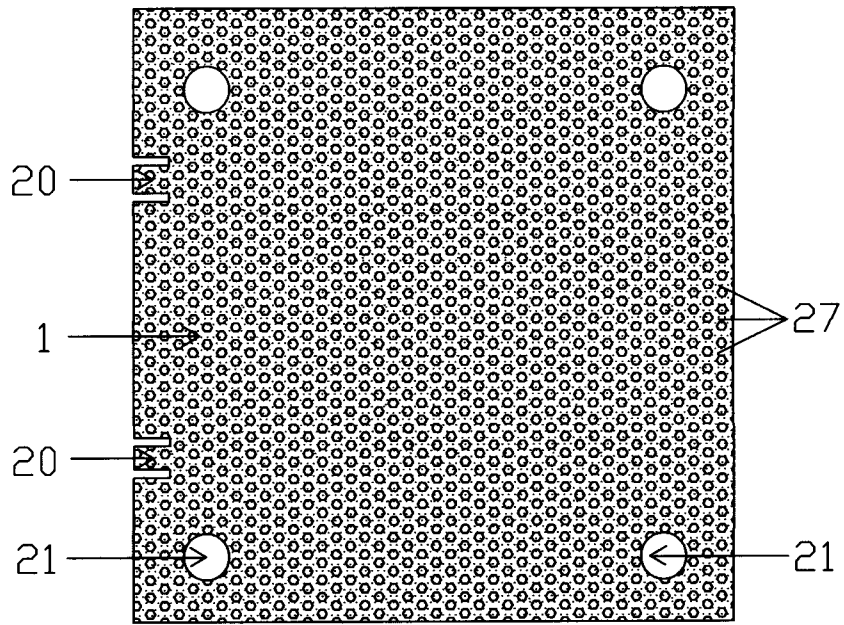
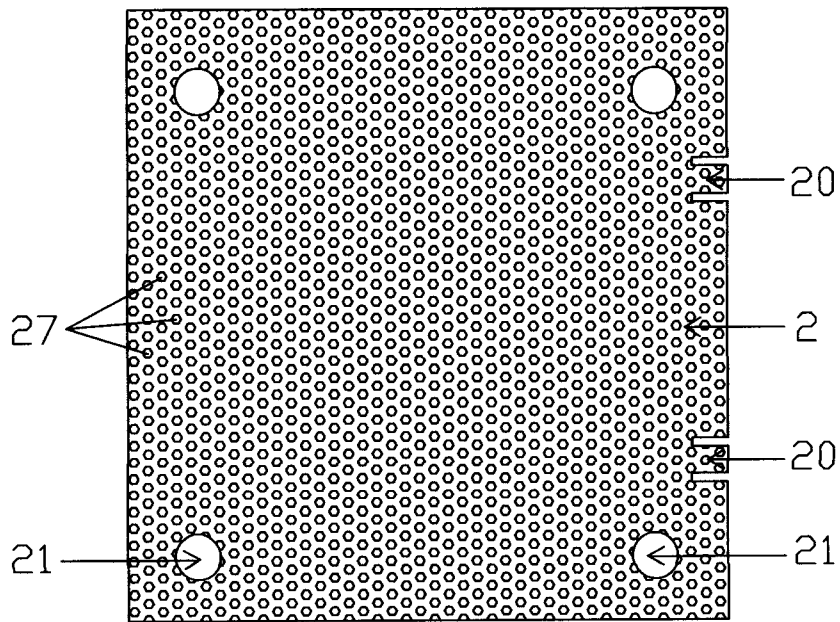


Figure 2



Anode



Cathode

Figure 3

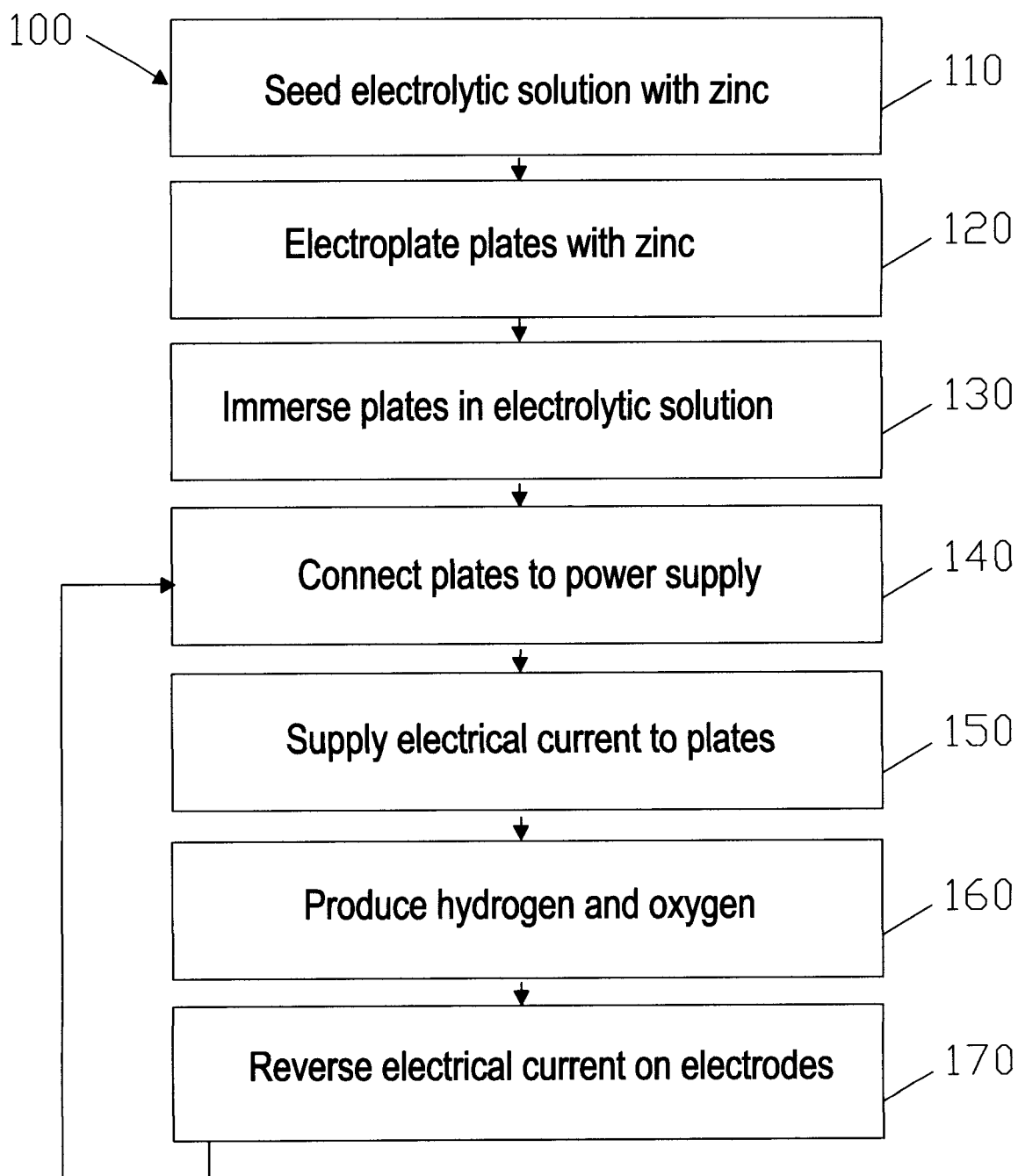


Figure 4

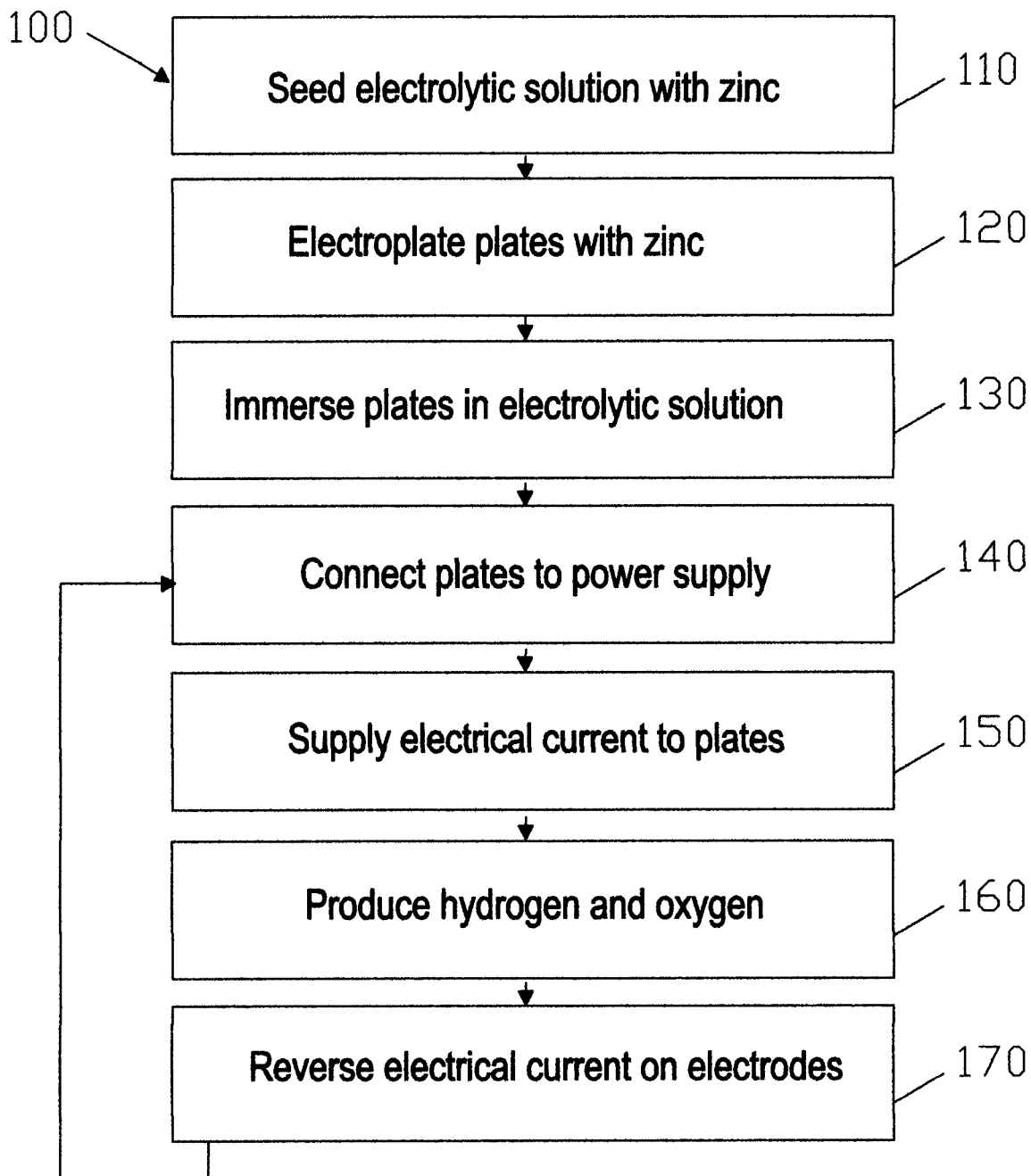


Figure 4