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 (54) Title: CATALYST MANUFACTURE

(57) **Abrégé/Abstract:**

A method is described for lowering the nitrate content in a cobalt catalyst precursor formed by impregnating a catalyst support with cobalt nitrate, comprising calcining the impregnated support in air to effect partial denitrification and subsequently heating the calcined impregnated support to a temperature below 2500C in the presence of a gas mixture comprising 0.1-10% hydrogen by volume in an inert gas.



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(54) Title: CATALYST MANUFACTURE

(57) Abstract: A method is described for lowering the nitrate content in a cobalt catalyst precursor formed by impregnating a catalyst support with cobalt nitrate, comprising calcining the impregnated support in air to effect partial denitrification and subsequently heating the calcined impregnated support to a temperature below 2500C in the presence of a gas mixture comprising 0.1-10% hydrogen by volume in an inert gas.



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Catalyst Manufacture

This invention relates to methods for manufacturing catalysts and in particular to the manufacture of cobalt catalysts from cobalt-nitrate-impregnated catalyst supports.

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Preparation of supported cobalt catalysts suitable for hydrogenation reactions or the Fischer-Tropsch synthesis of hydrocarbons has typically been by impregnation of soluble cobalt compounds into 'pre-formed' oxidic support materials or by precipitation of cobalt compounds from solution in the presence of support powders or extrudates, followed by a heating step in air and then, prior to use, activation of the catalyst by reduction of the resulting cobalt compounds in the catalyst precursors to elemental, or 'zero-valent' form typically using a hydrogen-containing gas stream. The heating in air converts at least some of the cobalt compounds to cobalt oxide, Co_3O_4 . The subsequent reduction with hydrogen converts the Co_3O_4 to cobalt monoxide, CoO , and thence the catalytically active cobalt metal.

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Impregnation methods typically rely on cobalt nitrate as it is relatively easy to manufacture at low cost. It is necessary, however, to reduce the residual nitrate (NO_3) level in the catalysts to very low levels to prevent emissions of nitrogen-oxide (NO_x) gases to the environment during subsequent processing. Whereas NO_x abatement technology, e.g. NO_x scrubbing, is commonly provided for calcination of nitrate-containing catalyst precursors, it is usually absent from catalyst reduction equipment. Furthermore, installing NO_x abatement technology is generally not practical where catalyst precursors are to be reduced to the active form in-situ, e.g. in a hydrogenation or Fischer-Tropsch reactor. However, we have found that in order to reduce the nitrate content of the catalyst precursor to acceptably low levels in the final catalyst precursor during the heating step in air, it is necessary to heat the precursor in air to temperatures in excess of 500°C . Prolonged heating of the catalyst precursor at these high temperature has been found to reduce the resulting cobalt surface area of the subsequently reduced catalysts, possibly as a result of increased support-metal interactions leading to undesired formation of spinel or other complex oxides. For example, heating cobalt compounds supported on alumina in air can increase cobalt aluminate formation. In the subsequent catalyst activation, cobalt aluminate is more resistant to reduction with hydrogen than cobalt oxide, requiring prolonged reduction times or increased temperatures. Both of these can lead to reduced cobalt surface areas in the resulting catalysts. As cobalt surface area has been found to be proportional to catalyst activity, a method for the preparation of the catalyst precursor at lower temperatures, but which also reduces nitrate levels to low levels is desirable.

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We have developed a method that overcomes these problems.

Accordingly the invention provides a method for lowering the nitrate content in a cobalt catalyst precursor formed by impregnating a catalyst support with cobalt nitrate, comprising calcining the impregnated support in air to effect partial denitrification and subsequently heating the calcined impregnated support to a temperature below 250°C in the presence of a gas mixture comprising 0.1-10% hydrogen by volume in an inert gas.

The method reduces nitrate content of the precursor to low levels without causing any reduction of the cobalt to elemental, or 'zero-valent', metallic form. By "without causing any reduction" we mean that the cobalt is not substantially reduced to metallic form, i.e. <1%, preferably <0.5%, more preferably <0.05% of the cobalt is reduced. Consequently the precursor is safe to handle. Nor is there we believe any conversion of Co_3O_4 to CoO and so possible re-oxidation of CoO is avoided.

The method of the present invention is performed on an impregnated catalyst support. Impregnation methods for producing cobalt catalysts are well known and generally comprise combining a catalyst support with a solution of cobalt nitrate, e.g. cobalt (II) nitrate hexahydrate at a suitable concentration. For example, an incipient wetness technique may be used whereby sufficient cobalt solution to fill up the pores of the support material added to the catalyst support. Alternatively larger amounts of cobalt solution may be used if desired. Whereas a number of solvents may be used such as water, alcohols, ketones or mixtures of these, preferably the support has been impregnated using aqueous solutions of cobalt nitrate. Single or multiple impregnations may be performed to achieve a desired cobalt level in the catalyst precursor.

The support material may comprise alumina, silica (including both synthetic silica and naturally occurring forms of silica such as kieselguhr), alumino-silicate, titania, zirconia, carbon or other suitable support such as zinc oxide. The support may comprise one or more of these, for example the support may be an aluminosilicate, a titania-coated silica or alumina, or a zirconia-coated silica or alumina. Preferably an alumina support is present, which is most preferably a transition alumina, so that preferred catalysts according to the invention comprise a cobalt species on a transition alumina support. Suitable transition alumina may be of the gamma-alumina group, for example a eta-alumina or chi-alumina. These materials may be formed by calcination of aluminium hydroxides at 400 to 750°C and generally have a BET surface area in the range 150 to 400 m^2/g . Alternatively, the transition alumina may be of the delta-alumina group which includes the high temperature forms such as delta- and theta- aluminas which may be formed by heating a gamma group alumina to a temperature above about 800°C. The delta-group aluminas generally have a BET surface area in the range 50 to 150 m^2/g . The transition aluminas contain less than 0.5 mole of water per mole of Al_2O_3 , the actual amount of

water depending on the temperature to which they have been heated. Alternatively, the catalyst support may comprise alpha-alumina.

The catalyst of the invention is particularly suitable for use in Fischer-Tropsch (F-T) hydrocarbon synthesis and the supports preferred for cobalt catalysts for use in known cobalt
5 F-T catalysts may be advantageously used for the catalysts of the present invention.

The support may be in the form of a powder or of a shaped unit such as a granule, tablet or extrudate. Shaped units may be in the form of elongated cylinders, spheres, lobed or fluted cylinders or irregularly shaped particles, all of which are known in the art of catalyst
10 manufacture. Alternatively the support may be in the form of a coating upon a structure such as a honeycomb support, monolith etc. Support materials may themselves contain promoters, stabilisers or other materials such as binders and may be treated prior to use in the process of the invention, e.g. by spray drying and/or calcining.

15 A suitable powder for the catalyst support generally has a surface-weighted mean diameter $D[3,2]$ in the range 1 to 200 μm . In certain applications such as for catalysts intended for use in slurry reactions, it is advantageous to use very fine particles which have a surface-weighted mean diameter $D[3,2]$ on average, in the range from 1 to 20 μm , e.g. 1 to 10 μm . For other applications e.g. as a catalyst for reactions carried out in a fluidised bed, it may be desirable to
20 use larger particle sizes, preferably in the range 50 to 150 μm . The term surface-weighted mean diameter $D[3,2]$, otherwise termed the Sauter mean diameter, is defined by M. Alderliesten in the paper "A Nomenclature for Mean Particle Diameters"; Anal. Proc., vol 21, May 1984, pages 167-172, and is calculated from the particle size analysis which may conveniently be effected by laser diffraction for example using a Malvern Mastersizer.

25 If desired, the impregnated support may be pre-dried to remove solvent prior to calcination and treatment with the hydrogen/inert gas mixture. The pre-drying step may be performed at 20-120°C, preferably 95-110°C, in air or under an inert gas such as nitrogen, or in a vacuum oven.

30 The cobalt nitrate impregnated catalyst precursor is heated in air to effect partial denitrication. The heating step, also termed herein as calcination, causes a physiochemical change in the impregnated support whereby the cobalt nitrate is at least partially decomposed to cobalt oxide, Co_3O_4 , with evolution of NO_x gases. The calcination temperature is preferably in the range 130 to 500°C but the maximum calcination temperature is preferably $\leq 400^\circ\text{C}$, more preferably
35 $\leq 350^\circ\text{C}$, most preferably $\leq 300^\circ\text{C}$, especially $\leq 250^\circ\text{C}$ to minimize cobalt-support interactions. The calcination time is preferably ≤ 24 , more preferably ≤ 16 , most preferably ≤ 8 , especially ≤ 6 hours. The calcination is most preferably performed $\leq 250^\circ\text{C}$ for ≤ 2 hours.

The cobalt content of the calcined impregnated support is preferably $\geq 5\%$ by weight, more preferably $\geq 15\%$ by weight. The upper limit may be 40% by weight.

5 The amount of nitrate present in the impregnated support depends in the amount of cobalt nitrate actually impregnated, i.e. is proportional to the cobalt content of the impregnated support. The calcination step preferably removes $>50\%$, more preferably $>75\%$, most preferably $>90\%$ of the nitrate originally present from the impregnated support. For example, where the cobalt content of the calcined impregnated support is about 20% wt, we have found that the residual nitrate content of the calcined impregnated support, expressed as NO_3 by
10 weight, prior to treatment with the hydrogen-containing gas is $>1\%$ wt, often 3 to 7% wt depending on the calcination conditions. (At 20% wt Co in the reduced catalyst, the impregnated support prior to the calcination step will have comprised about 30% wt NO_3). In the present invention, calcination is preferably operated such that the catalyst precursor contains nitrate levels in the range 1-10%, more preferably 1-5%. Residual nitrate levels $>1\%$
15 wt pose a considerable problem during the subsequent reduction stage, in particular where the reduction is carried out in-situ in a hydrogenation or Fischer-Tropsch reactor.

Methods for measuring the nitrate content include reduction to ammonia using Devarda's alloy followed by distillation and titration, or dissolution into water and measurement with a calibrated
20 ion-selective electrode. The preferred method is measurement by calibrated ion selective electrode.

To reduce the nitrate levels to tolerable levels, the calcined impregnated support, after cooling, is then heated to a temperature below 250°C , preferably below 225°C , especially $\leq 200^\circ\text{C}$ in
25 the presence of a gas mixture comprising 0.1-10% hydrogen by volume in an inert gas, to effect further denitrification of the catalyst support. Preferably the temperature to which the calcined impregnated support is heated is $>50^\circ\text{C}$, more preferably $>100^\circ\text{C}$, especially $\geq 140^\circ\text{C}$. The inert gas may be any suitable inert gas such as nitrogen, helium and argon and one or more inert gases may be used. Preferably the inert gas is nitrogen. The hydrogen
30 concentration in the inert gas is 0.1-10% by volume, preferably 1-5% by volume. It is possible to increase or reduce the hydrogen concentration in the inert gas within these limits during the denitrification step. The gas pressure during denitrification may suitably be from 1-50, preferably 1-20, more preferably 1-10 bar abs. The time for which the support is heated in the presence of the gas mixture will depend on a number of factors including the cobalt (and hence
35 nitrate) content. The denitrification process may be operated for ≤ 24 hours, preferably ≤ 16 hours, more preferably < 8 hours, especially ≤ 6 hours. Most preferably the catalyst precursor is heated at $140\text{-}225^\circ\text{C}$ for 1-8 hours, especially to about 200°C for 1-3 hours. Preferably a stream of the gas mixture is passed over and/or through the calcined impregnated support. The gas hourly space velocity (GHSV) of the hydrogen/inert gas stream may be $50\text{-}4000\text{hr}^{-1}$,

but is preferably $50-1000\text{hr}^{-1}$, more preferably $100-500\text{hr}^{-1}$ at normal temperature and pressure (NTP). Under these conditions essentially no reduction of the cobalt oxide takes place. Premature reduction is undesirable as it would pose handling problems for the resulting catalyst precursor.

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The pre-drying and/or denitrification may be carried out batch-wise or continuously, depending on the availability of process equipment and/or scale of operation.

10 The treatment of the calcined impregnated support with the hydrogen-containing gas preferably further reduces the nitrate content by $> 50\%$, more preferably $>75\%$. For example, a calcined impregnated support having a nitrate content of $>1\%$ wt preferably has $< 0.5\%$ wt, preferably $< 0.25\%$ wt, most preferably $< 0.1\%$ wt residual nitrate content after treatment with the $0.1 - 10\%$ vol hydrogen-containing gas.

15 The catalyst precursor may in addition to cobalt, further comprise one or more suitable additives or promoters useful in hydrogenation reactions and/or Fischer-Tropsch catalysis. For example, the catalysts may comprise one or more additives that alter the physical properties and/or promoters that effect the reducibility or activity or selectivity of the catalysts. Suitable additives are selected from compounds of metals selected from molybdenum (Mo), copper
20 (Cu), iron (Fe), manganese (Mn), titanium (Ti), zirconium (Zr), lanthanum (La), cerium (Ce), chromium (Cr), magnesium (Mg) or zinc (Zn). Suitable promoters include rhodium (Rh), iridium (Ir), ruthenium (Ru), rhenium (Re), nickel (Ni), platinum (Pt) and palladium (Pd). Preferably one or more promoters selected from Ni, Pt, Pd, Ir, Re or Ru are included in the catalyst precursor. Additives and/or promoters may be incorporated into the catalyst precursor by use of suitable
25 compounds such as acids, e.g. perhenic acid, chloroplatinic acid, metal salts, e.g. metal nitrates or metal acetates, or suitable metal-organic compounds, such as metal alkoxides or metal acetylacetonates. Typical amounts of promoters are $0.1 - 10\%$ metal by weight on cobalt. If desired, the compounds of additives and/or promoters may be added in suitable amounts to the cobalt nitrate solutions. Alternatively, they may be combined with the catalyst
30 precursor before or after drying/denitrification.

To render the catalyst catalytically active for Fischer-Tropsch or hydrogenation reactions, at least a portion of the cobalt oxide may be reduced to the metal. Reduction is preferably performed using hydrogen-containing gasses at elevated temperature.

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Before the reduction step, the catalyst precursor may, if desired, be formed into shaped units suitable for the process for which the catalyst is intended, using methods known to those skilled in the art.

Reduction may be performed by passing a hydrogen-containing gas such as hydrogen, synthesis gas or a mixture of hydrogen with nitrogen or other inert gas over the oxidic composition at elevated temperature, for example by passing the hydrogen-containing gas over the catalyst precursor at temperatures in the range 300-600°C for between 1 and 16 hours, preferably 1 – 8 hours. Preferably the reducing gas comprises hydrogen at >25% vol, more preferably >50% vol, most preferably >75%, especially >90% vol hydrogen. Reduction may be performed at ambient pressure or increased pressure, i.e. the pressure of the reducing gas may suitably be from 1-50, preferably 1-20, more preferably 1-10 bar abs. Higher pressures >10 bar abs may be more appropriate where the reduction is performed in-situ.

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Catalysts in the reduced state can be difficult to handle as they can react spontaneously with oxygen in air, which can lead to undesirable self-heating and loss of activity. Consequently reduced catalysts suitable for hydrogenation reactions are preferably passivated following reduction with an oxygen-containing gas, often air or oxygen in carbon dioxide and/or nitrogen. For catalysts suitable for Fischer-Tropsch processes, passivation is not preferred and the reduced catalyst is preferably protected by encapsulation of the reduced catalyst particles with a suitable barrier coating. In the case of a Fischer-Tropsch catalyst, this may suitably be a FT-hydrocarbon wax. Alternatively, the catalyst can be provided in the oxidic unreduced state and reduced in-situ with a hydrogen-containing gas.

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Whichever route is chosen, the cobalt catalysts prepared from precursors obtained by the method of the present invention provide high metal surface areas per gram of reduced metal. For example, the cobalt catalyst precursors, when reduced by hydrogen at 425°C, preferably have a cobalt surface area of $\geq 20 \text{ m}^2/\text{g}$ of cobalt as measured by H_2 chemisorption at 150°C. More preferably the cobalt surface area is $\geq 30 \text{ m}^2/\text{g}$ cobalt and most preferably $\geq 40 \text{ m}^2/\text{g}$ cobalt. Preferably, in order to achieve a suitable catalyst volume in hydrogenation and or Fischer-Tropsch processes, the catalysts have a cobalt surface area/g catalyst $\geq 10 \text{ m}^2/\text{g}$ catalyst, more preferably $\geq 15 \text{ m}^2/\text{g}$ catalyst.

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The cobalt surface area may be determined by H_2 chemisorption. A preferred method, used in the Examples, is as follows; Approximately 0.2 to 0.5 g of sample material, e.g. catalyst precursor, is firstly degassed and dried by heating to 140°C at 10°C/min in flowing helium and maintaining at 140°C for 60 minutes. The degassed and dried sample is then reduced by heating it from 140°C to 425°C at a rate of 3°C /min under a 50 ml/min flow of hydrogen and then maintaining the hydrogen flow at 425°C for 6 hours. Following this reduction, the sample is heated under vacuum to 450°C at 10°C/min and held under these conditions for 2 hours. The sample is then cooled to 150°C and maintained for a further 30 minutes under vacuum. The chemisorption analysis is then carried out at 150°C using pure hydrogen gas. An automatic analysis program is used to measure a full isotherm over the range 100 mm Hg up to

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760 mm Hg pressure of hydrogen. The analysis is carried out twice; the first measures the "total" hydrogen uptake (i.e. includes chemisorbed hydrogen and physisorbed hydrogen) and immediately following the first analysis the sample is put under vacuum (< 5mm Hg) for 30 mins. The analysis is then repeated to measure the physisorbed uptake. A linear regression is then applied to the "total" uptake data with extrapolation back to zero pressure to calculate the volume of gas chemisorbed (V).

Cobalt surface areas may then be calculated using the following equation;

$$\text{Co surface area} = (6.023 \times 10^{23} \times V \times \text{SF} \times A) / 22414$$

where V = uptake of H₂ in ml/g

SF = Stoichiometry factor (assumed 2 for H₂ chemisorption on Co)

A = area occupied by one atom of cobalt (assumed 0.0662 nm²)

This equation is described in the Operators Manual for the Micromeritics ASAP 2010 Chemi System V 2.01, Appendix C, Part No. 201-42808-01, October 1996.

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The catalysts may be used for hydrogenation reactions and for the Fischer-Tropsch synthesis of hydrocarbons.

Typical hydrogenation reactions include the hydrogenation of aldehydes and nitriles to alcohols and amines respectively, and the hydrogenation of cyclic aromatic compounds or unsaturated hydrocarbons. The catalysts of the present invention are particularly suitable for the hydrogenation of unsaturated organic compounds particularly oils, fats, fatty acids and fatty acid derivatives like nitriles. Such hydrogenation reactions are typically performed in a continuous or batch-wise manner by treating the compound to be hydrogenated with a hydrogen-containing gas under pressure in an autoclave at ambient or elevated temperature in the presence of the cobalt-catalyst, for example the hydrogenation may be carried out with hydrogen at 80-250°C and a pressure in the range 0.1- 5.0 x 10⁶ Pa.

The Fischer-Tropsch synthesis of hydrocarbons with cobalt catalysts is well established. The Fischer-Tropsch synthesis converts a mixture of carbon monoxide and hydrogen to hydrocarbons. The mixture of carbon monoxide and hydrogen is typically a synthesis gas having a hydrogen: carbon monoxide ratio in the range 1.7-2.5:1. The reaction may be performed in a continuous or batch process using one or more stirred slurry-phase reactors, bubble-column reactors, loop reactors or fluidised bed reactors. The process may be operated at pressures in the range 0.1-10Mpa and temperatures in the range 150-350°C. The gas-hourly-space velocity (GHSV) for continuous operation is in the range 100-25000hr⁻¹. The catalysts of the present invention are of particular utility because of their high cobalt surface areas/g catalyst.

The invention will now be further described by reference to the following Examples.

In each of the Examples an impregnated support was prepared as follows; a gamma alumina support (Puralox 20/190 from Sasol Condea) was impregnated using a solution comprising
5 125.91 g of $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$, 0.91 g NH_4ReO_4 and 29.4 g of demineralised water per 100 g support, to form a damp cake. This was dried on trays in an oven at 110°C for 4 hours at a bed depth of 25 mm.

In each Example, nitrate (NO_3) contents were measured as follows; a known weight of sample
10 is mixed with 50 ml demineralised water in a beaker, covered and heated to boiling for 1 minute. This is then cooled and 10 ml of 2M ammonium sulphate solution is added. This is then made up to 100 ml in a volumetric flask and the sample left for 10 minutes to settle. 75 ml are then decanted to a stirred beaker and the conductivity measured using a nitrate ion selective electrode. The conductivity is then correlated to a calibration curve prepared using
15 known nitrate levels to obtain the nitrate content of the sample.

In the Examples, the catalyst precursors were reduced at 425°C using hydrogen and their cobalt surface areas determined by hydrogen chemisorption at 150°C according to the method described above.

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The cobalt contents of the reduced catalysts were all 20% by weight.

Comparative Example 1 (not according to the invention)

The dried material was calcined in air at a bed depth of 150 mm. The temperature of the
25 powder during calcination was raised to 370°C over 2 hours, then to 400°C over 1.5 hours before being lowered to 270°C over 1.5 hours.

The NO_3 content of the calcined catalyst precursor was 0.92% by weight. The cobalt surface area of the reduced catalyst precursor was $8.2 \text{ m}^2/\text{g}$ reduced catalyst.

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Thus calcination alone results in a low cobalt surface area catalyst and the nitrate content of the precursor is still ca 1% wt.

Comparative Example 2 (not according to the invention)

35 The dried material was calcined in air at a bed depth of 50 mm. The temperature of the powder during calcination was raised to 180°C over 18 minutes, then to 200°C over 10 minutes, then held at 200°C for 10 minutes before being lowered to 130°C over 20 minutes.

The NO₃ content of the calcined material was 1.8% by weight. The cobalt surface area of the reduced calcined material was 11.8 m²/g reduced catalyst. This calcined material was further calcined in air using a fluidised bed using 0.4 Nm³/hr air per kg oxide (SV = 400hr⁻¹). The temperature was raised from ambient to 200°C over 1.5 hours and held at 200°C for 2 hours.

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The NO₃ content of this calcined catalyst precursor was 1.6% by weight and the cobalt surface area of the reduced material was almost unchanged at 11.6 m²/g reduced catalyst. Thus lower temperature calcination, even in two steps, gives an increase in catalyst cobalt surface area but the level of residual NO₃ in the precursor would pose an environmental problem during

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reduction of the catalyst.

Example 3 (according to the invention)

The dried material was calcined at a bed depth of 50 mm. The temperature of the powder during calcination was raised to 180°C over 18 minutes, then to 200°C over 10 minutes, then held at 200°C for 10 minutes before being lowered to 130°C over 20 minutes.

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The NO₃ content of the calcined material was 1.8% by weight. The cobalt surface area of the reduced calcined material was 11.8 m²/g reduced catalyst. This calcined material was further heated in a fluidised bed using 0.2 Nm³/hr of 5% v/v hydrogen in nitrogen per kg oxide (SV = 200 hr⁻¹). The temperature was raised from ambient to 200°C over 1.5 hours and held at 200°C for 2 hours.

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The cobalt surface area of the reduced catalyst precursor prepared using this material prepared according to the invention was 11.7 m²/g reduced catalyst, but the NO₃ content of the catalyst precursor was reduced markedly to 0.08% by weight. Thus the present invention provides a method whereby the nitrate content of catalyst precursors may be reduced without sacrificing cobalt surface area.

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Claims.

1. A method for preparing a cobalt catalyst precursor by (i) impregnating a catalyst support with cobalt nitrate, (ii) calcining the impregnated support in air to effect partial denitrication and form Co_3O_4 and (iii) subsequently heating the calcined impregnated support comprising Co_3O_4 to a temperature below 225°C in the presence of a gas mixture comprising 0.1-10% hydrogen by volume in an inert gas to form a catalyst precursor without causing any reduction of the cobalt to metallic form.
2. A method according to claim 1 wherein the calcination is performed at a temperature in the range $130 - 500^\circ\text{C}$.
3. A method according to claim 2 wherein the maximum calcination temperature is $\leq 400^\circ\text{C}$.
4. A method according to any one of claims 1 to 3 wherein the calcined impregnated support is heated to a temperature above 50°C .
5. A method according to any one of claims 1 to 4 wherein the inert gas is nitrogen.
6. A method according to any one of claims 1 to 5 wherein the hydrogen concentration in the inert gas is 1-5% by volume.
7. A method according to any one of claims 1 to 6 wherein the pressure of the hydrogen-containing gas mixture is 1-50 bar abs.
8. A method according to any one of claims 1 to 7 wherein the cobalt content of the calcined impregnated support is between 5 and 40% by weight.
9. A method according to any one of claims 1 to 8 wherein the catalyst precursor, in addition to cobalt, further comprises one or more suitable additives or promoters useful in hydrogenation reactions and/or Fischer-Tropsch catalysis.
10. A method for preparing a cobalt catalyst comprising the steps;
 - (i) preparing a catalyst precursor according to the method of claims 1 to 9, and
 - (ii) reducing the catalyst precursor in a stream of hydrogen-containing gas between 300 and 600°C .