

(19) World Intellectual Property Organization
International Bureau



(43) International Publication Date
7 July 2011 (07.07.2011)

PCT

(10) International Publication Number
WO 2011/081727 A2

(51) International Patent Classification:

F01N 3/08 (2006.01) *B01J 21/06* (2006.01)
B01D 53/94 (2006.01) *F01N 3/20* (2006.01)

(21) International Application Number:

PCT/US2010/056485

(22) International Filing Date:

12 November 2010 (12.11.2010)

(25) Filing Language:

English

(26) Publication Language:

English

(30) Priority Data:

12/638,166 15 December 2009 (15.12.2009) US

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(81) Designated States (unless otherwise indicated, for every

kind of national protection available): AE, AG, AL, AM, AO, AT, AU, AZ, BA, BB, BG, BH, BR, BW, BY, BZ, CA, CH, CL, CN, CO, CR, CU, CZ, DE, DK, DM, DO, DZ, EC, EE, EG, ES, FI, GB, GD, GE, GH, GM, GT, HN, HR, HU, ID, IL, IN, IS, JP, KE, KG, KM, KN, KP, KR, KZ, LA, LC, LK, LR, LS, LT, LU, LY, MA, MD, ME, MG, MK, MN, MW, MX, MY, MZ, NA, NG, NI, NO, NZ, OM, PE, PG, PH, PL, PT, RO, RS, RU, SC, SD, SE, SG, SK, SL, SM, ST, SV, SY, TH, TJ, TM, TN, TR, TT, TZ, UA, UG, US, UZ, VC, VN, ZA, ZM, ZW.

(84) Designated States (unless otherwise indicated, for every

kind of regional protection available): ARIPO (BW, GH, GM, KE, LR, LS, MW, MZ, NA, SD, SL, SZ, TZ, UG, ZM, ZW), Eurasian (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European (AL, AT, BE, BG, CH, CY, CZ, DE, DK, EE, ES, FI, FR, GB, GR, HR, HU, IE, IS, IT, LT, LU, LV, MC, MK, MT, NL, NO, PL, PT, RO, RS, SE, SI, SK, SM, TR), OAPI (BF, BJ, CF, CG, CI, CM, GA, GN, GQ, GW, ML, MR, NE, SN, TD, TG).

Published:

— without international search report and to be republished upon receipt of that report (Rule 48.2(g))

(54) Title: SELECTIVE CATALYTIC REDUCTION SYSTEM FOR CAPTURE OF VOLATILIZED COMPOUNDS

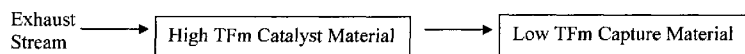


Figure 1

(57) Abstract: An apparatus and method for treating diesel exhaust gases are described. The system consists of two functionalities, the first being a selective catalytic reduction (SCR) catalyst system and the second being a capture material for capturing catalyst components that have appreciable volatility under extreme exposure conditions. The SCR catalyst component is typically based on a majority phase of titania, with added minority-phase catalyst components comprising of one or more of the oxides of vanadium, silicon, tungsten, molybdenum, iron, cerium, phosphorous, copper and/or manganese vanadia. The capture material typically comprises a majority phase of high surface area oxides such as silica-stabilized titania, alumina, or stabilized alumina, for example, wherein the capture material maintains a low total fractional monolayer coverage of minority phase oxides for the duration of the extreme exposure. The method involves treatment of hot exhaust streams by both the catalyst material and capture material, wherein the capture material can be in a mixture with the catalyst material, or can be located downstream thereof, or both, but still be maintained at the extreme temperatures. Volatile catalyst components such as vanadia and tungsta are thus removed from the vapor phase of the exhaust gas.



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SELECTIVE CATALYTIC REDUCTION SYSTEM FOR CAPTURE OF VOLATILIZED COMPOUNDS

BACKGROUND

[0001] The selective catalytic reduction (SCR) of nitrogen oxides produced in combustion engines, with reductants such as urea and ammonia, is an industrially significant catalytic process. Vanadia-based SCR catalysts that utilize titania catalyst supports are approved for use for on-road mobile applications in Europe on Heavy-Duty Diesel trucks, and these catalyst are highly active and show excellent tolerance to fuels that contain sulfur. However, vanadia-based catalysts are not approved by the EPA for on-road use in the U.S. or in Japan. This lack of approval stems from the concern over release of vanadia into the environment and the potential toxicity that might arise from exposure to vanadia emitted from the tailpipe. One possible mechanism that potentially might cause a loss of vanadia from the catalyst is vaporization of the metal oxide or hydroxide at high temperature in the stream of hot exhaust gases.

[0002] Furthermore, newer regulations for soot and NO_x that will begin to be imposed as early as 2010 (e.g., Euro VI and US 2010 regulations) may necessitate the use of both a diesel particulate filter (DPF) in tandem with an SCR catalyst. In one configuration (U.S. Patent No. 7,498,010) the SCR catalyst is downstream from the DPF. If no remedial action is taken, the collection of soot on the DPF, will ultimately plug the channels for exhaust gas flow, and can cause unacceptable pressure-drop arises across the device. In order to avoid this situation, the soot is removed either continuously or sporadically by combustion. Since combustion is an exothermic process, it is associated with a rise in the temperature of the device that is transmitted to the exhaust gases, and the temperature rise depends on the amount of soot collected as well as the temperature of the exhaust gas upstream of the DPF. These high temperature exhaust gases, which can approach 750°C and higher, subsequently will pass over the SCR catalyst. Thus, there has been much emphasis recently on improving the thermal stability of the SCR catalyst, both for vanadia-based catalysts as well as Cu-, Fe- and other base-metal catalysts. It is generally accepted that the catalyst must be stable to temperatures of up to 800°C for short periods of time. In order to test the durability of catalyst

formulations, it is necessary to develop tests that simulate real-world exposure conditions. Ford researchers[1] have developed an accelerated aging protocol for SCR catalysts that simulates on-road conditions over a span of 120,000 miles. This test involves exposing the catalyst to a reactant gas stream that includes water (5 vol%), for a time of 64 hr at 670°C at relatively high gas flow rate (gas hourly space velocity, GHSV = 30,000 hr⁻¹). These conditions of time and temperature are used as a reference point for the method of the presently disclosed and claimed inventive concepts.

[0003] Concern over the volatility of vanadia at high temperatures, e.g., when the SCR catalyst is located downstream of the DPF, is thus an issue that may limit the available market for vanadia-based mobile SCR catalysts and is a key consideration in catalyst development. There has thus remained in the art a need to be able to evaluate the degree of vanadia volatilization from SCR catalysts. Further, there has remained a need in the art for a deNO_x selective catalytic reduction catalyst system which demonstrates zero vanadia loss downstream thereof. It is an object of the presently disclosed and claimed inventive concepts to address these shortcomings of the prior art.

SUMMARY OF THE DISCLOSURE

[0004] The present disclosure describes a selective catalytic reduction catalyst system for the capture of volatilized compounds such as vanadium and tungsten compounds, and compositions and processes for the capture of volatilized compounds such as vanadium and tungsten compounds in such a selective catalytic reduction catalyst system, for example in an emission control system of a diesel engine. The system typically comprises two functionalities, the first being a selective catalytic reduction (SCR) catalyst system and the second being a capture material for capturing catalyst components that have appreciable volatility under extreme exposure conditions. The SCR catalyst component is typically based on a majority phase of titania, with added minority-phase catalyst components comprising of one or more of the oxides of vanadium, silicon, tungsten, molybdenum, iron, cerium, phosphorous, copper and/or manganese vanadia. The capture material typically comprises a majority phase of high surface area oxides such as silica-stabilized titania, alumina, or stabilized alumina, for example, wherein the capture material maintains a low total fractional monolayer coverage of minority phase oxides for the duration of the extreme exposure. The method involves treatment of hot exhaust streams

by both the catalyst material and capture material, wherein the capture material can be in a mixture with the catalyst material, or can be located downstream thereof, or both, but still be maintained at the extreme temperatures. Volatile catalyst components such as vanadia and tungsta are thus removed from the vapor phase of the exhaust gas.

[0005] In one aspect of the disclosure, it is contemplated that stable, high surface area oxide supports, including but not limited to, silica-stabilized titania and alumina, can be employed to capture these volatile components, when said components are present at low density on the support surface of the capture bed after their capture. Further, in regard to certain embodiments, it has also been surprisingly found that when the normally volatile components are present at low density on the catalyst surface prior to exposure to extreme conditions, they also exhibit diminished volatility at the extreme conditions, so that such "low density" catalysts can also be employed to reduce or eliminate catalyst component volatility. Thus, the volatile components, once captured, preferably will not be substantially released again into the exhaust phase. Therefore, in one aspect the system and method provides high stability, high surface area inorganic oxide supports that can be used in a configuration wherein the stable, high surface area supports are either mixed with, or are positioned downstream from a lower stability, V-based SCR catalyst material in which the catalyst components are present at relatively high surface density, wherein the former stable, high surface area supports capture and remove volatile compounds from the vapor phase even at the same temperature which is present in the lower stability and higher surface density catalyst material from which the compounds were volatilized.

[0006] The present summary is not intended to be an exhaustive or complete summary of the disclosure but is only intended to identify various notable aspects thereof. Other aspects of the disclosure not specifically noted above will become evident upon consideration of the description below.

BRIEF DESCRIPTION OF THE DRAWINGS

[0007] Figure 1 is a schematic representation of one embodiment of a selective catalytic reduction catalyst system of the presently disclosed and claimed inventive concepts wherein the catalyst material is positioned upstream of the capture material.

[0010] Figure 2 is a schematic representation of another embodiment of a

selective catalytic reduction catalyst system of the presently disclosed and claimed inventive concepts wherein the catalyst material and capture material are combined in a mixture.

[0011] Figure 3 is a graph showing the relationship between the Total Fractional Monolayer values and the vanadium volatility

[0012] Figure 4 is a graph showing the relationship between the Total Fractional Monolayer values and the tungsten volatility.

DETAILED DESCRIPTION

[0013] In the course of investigation of the vaporization of catalyst components (supported on oxide supports) during high-temperature, accelerated aging tests, it has surprisingly been found that the catalyst components exhibit varying degrees of volatility depending on the catalyst support, and the volatilities of the supported catalyst materials can be substantially different from volatilities of the bulk oxides. In the event that the supported catalyst shows non-negligible vaporization of the catalyst components, it is desirable to provide a means of capturing the volatilized components released into the vapor stream. One means of capturing such components that are volatile at high temperature is simply to allow them to condense at a lower temperature downstream of the catalyst bed. However, this approach is problematic because the volatile components may condense on unintended locations. Thus, it is even more desirable to provide a means for capturing the volatile components even at very high sustained temperatures such as might be encountered in the catalyst bed. The presently disclosed and claimed inventive concepts therefore are directed to a treatment system for diesel engine exhaust, and use therewith, which comprises a vanadia-based selective catalytic reduction catalyst (V-based SCR catalyst) system and which comprises a "capture-bed" either mixed with, or immediately downstream of the V-based SCR catalyst. The function of the capture bed is to capture and retain any volatile compounds even at elevated temperatures and gas flow conditions which approach those experienced in the catalytic portion of the catalyst/capture bed mixture, or in the upstream catalyst bed, so that the volatile components are thus removed from the vapor phase of the exhaust gas.

[0014] In one embodiment of the presently disclosed and claimed inventive concepts, where the catalyst material and capture bed material comprise a mixture, the v/v ratio of the catalyst material and capture bed material may be,

for example, in the range of 1:20 to 20:1, and more preferably be, for example, 1:10 to 10:1.

[0015] Further, where used herein, the terms "catalyst bed," "catalyst material," and "catalyst bed material" may be used interchangeably. Similarly, the terms "capture bed," "capture material," and "capture bed material" may be used interchangeably.

[0016] Where used herein the term "substantially all" means at least 90% of the material which is referred to, or more preferably to at least 95% of the material which is referred to, or more preferably to at least 96% of the material which is referred to, or more preferably to at least 97% of the material which is referred to, or more preferably to at least 98% of the material which is referred to, or more preferably to at least 99% of the material which is referred to.

[0017] It has presently been found that stable, high surface area oxide supports, including but not limited to, silica-stabilized titania or alumina, can be employed to capture the volatile components, when said components are present at low density on the support surface of the capture bed material after their capture. Further, in regard to certain embodiments, it has also been surprisingly found that when the normally volatile components are present at low density on the catalyst surface prior to exposure to extreme conditions, they also exhibit diminished volatility at the extreme conditions, so that such "low density" catalysts can also be employed to reduce or eliminate catalyst component volatility. Thus, a key aspect of the presently disclosed and claimed inventive concepts is to provide for high stability, high surface area inorganic oxide supports that can be used in a configuration wherein the stable, high surface area supports are either mixed with, or are positioned downstream from a lower stability, V-based SCR catalyst material in which the catalyst components are present at relatively high surface density, wherein the former stable, high surface area supports capture and remove volatile compounds from the vapor phase even at the same temperature which is present in the lower stability and higher surface density catalyst material from which the compounds were volatilized.

[0018] This configuration, shown in a schematic diagram in Fig. 1, can be achieved, in one exemplary embodiment by preparing wash-coated catalysts that are "zoned", that is, the lower stability, higher surface density V-based SCR catalyst is located in a position in front of the device, while the high stability,

high surface area low surface density inorganic oxide is located in a position that is towards the back of the device.

[0019] Another configuration, shown in a schematic diagram in Fig. 2, can be achieved in one exemplary embodiment by the co-extrusion (or other mode of mixing) of catalyst materials and capture bed materials. When the high stability, high surface area oxide is comprised of a composition that is a good support for V-based SCR catalysts, then the capture material may be an active catalyst even as it accumulates the volatile oxides from the less stable catalyst. Titania-based materials that are suitable to be used in this configuration as the highly stable, high surface area supports are those described in a recent patent application (U.S. Serial No. 12/533,414). It will be seen below that an example of a lower stability, higher surface density V-based SCR catalyst is 2 wt% vanadia supported on the commercially-available catalyst DT-52™ (DT-52™ itself is a material comprising about 10 wt% WO₃ and about 90 wt% TiO₂ (anatase form) and has a BET surface area of about 90 m²/g). This catalyst is nonetheless a highly active SCR catalyst under normal conditions, but exhibits considerable loss of surface area after exposure to extreme conditions. Thus, one embodiment of the presently disclosed and claimed inventive concepts comprises a mixture of a vanadia on DT-52™ catalyst material with the titania-based materials or catalysts described in the recent application (U.S. Serial No. 12/533,414). For example, the mixture can be achieved either in wash-coat or extrusion applications or any other suitable method for creating highly interspersed mixtures of particulate materials. In a second embodiment, the present invention comprises a system of a high surface area, high stability alumina (or any other suitable capture bed material) positioned downstream of the vanadia on DT-52™ catalyst (or any other appropriate catalytic material). For example, titania-based support materials which can be used herein include, but are not limited to, materials which comprises anatase titania particles comprising ≥ 85 wt% dry weight of TiO₂ and ≤ 10 wt% dry weight of SiO₂, wherein the SiO₂ is substantially in a low molecular weight and/or small nanoparticle form. The material may further comprise, for example, 3% to 10% WO₃ and may have a BET surface area of at least 80 m²/gm. The material may comprise $\geq 85\%$ dry weight of TiO₂, 3% - 10% of SiO₂, and 3% - 10% dry weight of WO₃, for example. The SiO₂ may be present at a fractional monolayer value of less than 1.0 before the material is sintered.

The small nanoparticle form of the SiO_2 may comprise a diameter of $< 5\text{nm}$. The low molecular weight form of the SiO_2 may comprise a MW of $< 100,000$. The SiO_2 may comprise silicon atoms which are substantially (e.g., $> 50\%$) in the Q^3 , Q^2 , Q^1 and Q^0 coordination environments. The SiO_2 may comprise patches which are substantially $\leq 5\text{ nm}$ deep after redistribution as seen by scanning electron microscopy or by transmission electron microscopy. The TiO_2 used may optionally not be prepared in the presence of urea. The material, in a preferred embodiment, comprises silica which is substantially in a low molecular weight form and/or a small nanoparticle form. By this is meant that over 50% of the silica is either in the low molecular weight form (MW $< 100,000$) or in a small nanoparticle form (diameter $< 5\text{ nm}$), or is a combination of both. In a more preferred version, the silica comprises $> 60\%$ low molecular weight forms and/or small nanoparticle forms. In a still more preferred version, the silica comprises $> 70\%$ low molecular weight and/or small nanoparticle forms. In a still more preferred version, the silica comprises $> 80\%$, and yet still more preferred, $> 90\%$, low molecular weight forms and/or small nanoparticle forms of silica. Furthermore, the low molecular weight and small nanoparticle forms of the material preferably have geometric surface areas of $> 450\text{ m}^2/\text{g}$.

[0020] Alternatively, the material comprises $\geq 85\%$ dry weight of TiO_2 , 5.0%-9.0% dry weight of SiO_2 , and 3.0%-7.0% dry weight of WO_3 . More particularly, the material comprises 87%-89% dry weight of TiO_3 , 7%-9% dry weight of SiO_2 , and 3%-5% dry weight of WO_3 . In one preferred embodiment the material comprises about 88% ($\pm 0.5\%$) dry weight TiO_2 , about 8% ($\pm 0.5\%$) dry weight SiO_2 , and about 4% ($\pm 0.5\%$) dry weight WO_3 . In one embodiment, the weight % of WO_3 is less than the weight % of SiO_2 . In one embodiment, the material has a fresh surface area of at least $80\text{ m}^2/\text{gm}$, and more preferably at least $100\text{ m}^2/\text{gm}$. In another embodiment, the material comprises $\geq 85\%$ dry weight of TiO_2 , 3.0%-8.0% dry weight of SiO_2 , and 4.0%-9.0% dry weight of WO_3 . More particularly, the material comprises $\geq 87\%$ dry weight of TiO_3 , 3%-6% dry weight of SiO_2 , and 4%-8% dry weight of WO_3 . In a preferred embodiment the material comprises about 90% ($\pm 0.5\%$) dry weight TiO_2 , about 4% ($\pm 0.5\%$) dry weight SiO_2 , and about 6% ($\pm 0.5\%$) dry weight WO_3 . In one

embodiment, the weight % of WO_3 is greater than the weight of % of SiO_2 . In one embodiment, the material has a fresh surface area of at least $80 \text{ m}^2/\text{gm}$, and more preferably at least $100 \text{ m}^2/\text{gm}$. In one embodiment, the TiO_2 component of the material used herein substantially comprises a surface area $<400 \text{ m}^2/\text{g}$ and a pore volume $<0.40 \text{ cm}^3/\text{g}$.

[0021] The embodiments of the presently disclosed and claimed inventive concepts thus enable the use of V-based SCR catalysts in configurations where the catalyst material and also the capture bed material, will be exposed to very high temperatures.

[0022] It is helpful to consider parameters that serve to define the capture materials, and one such parameter involves the surface density ρ_{surf} , (atoms/ nm^2) of catalyst components on a support material. The catalyst and capture materials of the presently disclosed and claimed inventive concepts are typically composed of a majority phase (mass fraction greater than about 0.7) and one or more minority phases (mass fraction less than about 0.3) wherein the minority phase is viewed as existing on the surface of the majority phase. For example, titania-based SCR catalysts typically contain a minority phase comprising vanadia (generally less than 5 %), tungsta (generally less than 15%), and optionally silica (generally less than 15%), that have been deposited on the surface of the titania (majority phase). When the minority phases are present at very low mass fraction, they can exist in an atomically dispersed state, bonded solely to the support material. By contrast, when the minority phases are present at higher mass fractions, they can begin to form bonds homogeneously and thus form one or more layers on the support material. In extreme cases, the minority phases may actually crystallize, forming e.g., bulk crystalline tungsta in its native monoclinic crystalline form, in intimate mixture with the support titania. In this state, it is possible that the minority phases develop chemical character more typical of that of the bulk phase of the minority oxide. In the particular cases of vanadia and tungsta, the bulk oxides may exhibit volatility under extreme conditions of temperature, water vapor, gas flow rate and time.

[0023] As can be ascertained from the chemical literature, the maximum amount of vanadia that can be maintained in a highly dispersed state on titania without vanadia crystal formation[2] is $7.1 \text{ V atoms}/\text{nm}^2$, and this density is assumed to represent monolayer coverage. Monolayer coverage of highly

dispersed tungsta on titania[3] is estimated to occur at a tungsta loading of 4.5 W/nm², and that for silica[4] is estimated to occur at a silica loading of 6 Si atoms/nm². These complete, or monolayer surface densities, $\rho_{\text{surf, monolayer, } i}$, for the highly dispersed minority phases can be used to further define the actual fractional coverage at the mass fraction, $f_{m, i}$, of the minority phases (i) in the real catalyst:

$$f_{m, i} = \rho_{\text{surf, } i} / \rho_{\text{surf, monolayer, } i}$$

[0024] In the case that there are multiple supported oxides (e.g., silica, tungsta and vanadia) in the real catalyst, the total fractional monolayer coverage Tf_m can then be defined as the sum of the $f_{m,i}$ for each of the minority-phase supported oxides. From the above definitions, it is apparent that the combination of high majority-phase support surface area and low minority-phase mass fraction results in very low fractional coverage for the component oxides. This combination is very desirable for the capture materials of the presently disclosed and claimed inventive concepts. Furthermore, since the object of the presently disclosed and claimed inventive concepts is to capture volatile oxides under conditions of extreme exposure, it follows that it is highly desirable that the surface area for the majority-phase oxide be stable and not substantially decrease as a result of the exposure conditions. The relevant surface area to be used in this definition is that which is measured after exposure of the capture or catalyst materials to the severe conditions that simulate lifetime exposure. Stable surface area thus means that there is minimal loss in surface area from the fresh state (before exposure) and the aged state (after exposure). The condition of low fractional coverage is also a desirable feature of the catalyst materials themselves, as the component volatility is surprisingly found to be minimal under that condition. It follows that it is also desirable for the catalyst material to retain appreciable surface area during exposure to harsh conditions.

[0025] The examples and embodiments herein and below refer to various titania and alumina materials which may comprise the majority phase of the catalyst support material or capture bed material of the invention. However, the majority phases of the catalyst support and capture bed materials are not to be limited to these and may also comprise, alone or in combination, other titanias, silica-stabilized titanias, aluminas (including, but not limited to, boehmite, gamma and alpha alumina), stabilized aluminas (for example, those stabilized by

lanthanum or other lanthanides), amorphous silicas, silica-aluminas, zeolites (including, but not limited to, faujasite, mordenite, ZSM-5 and Beta zeolite), and/or molecular sieves. In one embodiment, the minority phase of the capture material used in the selective catalytic reduction catalyst system of the presently disclosed and claimed inventive concepts maintains a total fractional monolayer coverage on the majority phase of about 5 or less under conditions encountered in accelerated aging tests developed by the Ford Co. [1] to simulate the lifetime (e.g., 120,000 miles) of on-road exposure to an SCR catalyst that is positioned downstream from a DPF. The Ford test conditions [1] are, for example, exposure to 670°C for 64 hr with a gas-hourly space velocity (GHSV) of 30,000 hr⁻¹ with 5 vol% water. Other test conditions which may be employed are 750°C for 4 hr and a GHSV of ~100,000 hr⁻¹ and 5 vol% water. For example, the materials referred to herein as MC-X5v1 and MC-X5v2 have T_{f_m} less than or equal to about 3 after these exposures. In another version illustrative of a material that does not meet the requirements for capture material of the present invention, DT-52™ w/2 wt% V₂O₅ has a T_{f_m} greater than about 3 after exposure at the relatively mild condition of 670°C for 4 hr with GHSV > 10,000 hr⁻¹ and ~5% H₂O. Preferably, the exposure conditions simulate the lifetime exposure of the catalysts under real-world conditions. Also, the degree of aging that occurs (and hence the final surface area and T_{f_m}) also depends on vanadia content. In the present disclosure the vanadia content of the catalysts are preferably in a range of between 0.5% and 5% vanadia, and more preferably in a range of between 1% and 3% vanadia.

[0026] EXAMPLES

[0027] The following experimental apparatus was designed to provide illustration of the presently disclosed and claimed inventive concepts. SCR catalyst samples consisting of catalyst components that include one or more of vanadia, tungsta and/or silica, were aged at elevated temperatures in a gas stream that contains H₂O and O₂, each at 5 vol %, and NO and NH₃, each at 500 ppm, with the balance consisting of N₂. This gas stream is a representative mixture that approximates the composition of gases under realistic conditions. The inorganic vapors generated from the catalyst samples were then captured on a "capture bed" located downstream of the catalyst but still in the hot zone of the furnace. For the approach to be successful, the volatile component vapors, in

this case, the oxides and hydroxides of vanadia and tungsta, must react rapidly and quantitatively with the capture material at the elevated temperatures of the test. If this condition is met, then the amount captured downstream also is an indirect measure of the vapor pressure oxide of interest. In the following discussion, the two quantities (amount captured and vapor pressure of volatile component) are used interchangeably.

[0028] Example 1- Mass Balance.

[0029] In this example, the use of a catalyst material comprising DT-52™ with 2 wt% vanadia (prepared by evaporation of an alkaline monoethanolamine/vanadia solution) was held in a position upstream of a capture bed, in tests to demonstrate that the volatile components from the catalyst sample could be quantitatively captured by the downstream bed. The DT-52™ support is commercially available from Millennium Inorganic Chemical, Inc. (MIC), and has a composition of 90% TiO₂, 10% WO₃. A gamma alumina (Alfa Aesar, aluminum oxide, gamma, catalyst support, high surface area, bimodal) was used as the capture bed material and was additionally calcined at 800°C for 6 hr in air, and had a nominal surface area of 200 m²/g. A small quantity, 0.2 g of -14/+24 mesh, of the alumina was placed in a reactor tube in a position downstream of the catalyst sample. The alumina capture bed was separated from an equivalent amount (0.2 g) of -14/+24 mesh of the catalyst sample by a short (< 1 cm) plug of quartz wool. A second, short plug of quartz wool downstream of the alumina capture bed was used to maintain the position of the alumina. DT-52™ with 2 wt% vanadia was chosen as the catalyst sample in this experiment because it is known in the art that this material does not have a high degree of stability at high temperatures. To confirm that fact, the surface area of the starting, vanadia-doped DT-52™ was 58 m²/g, while the surface area of the exposed and recovered material (as described below) was 12 m²/g. Hence the catalyst material underwent significant loss of surface area during the exposure. The catalyst material and alumina capture bed material were then exposed at 750°C for 1 hr with a total gas flow of 65 L/hr, and both the catalyst sample and capture bed sample were manually recovered for analysis. The temperature for exposure of 750°C was chosen instead of 670°C since the former causes a comparable amount of catalyst to occur in a shorter period of time so that the test could be shortened from 64 hr to 1-4 hr, while still providing representative results. This gas flow represents a GHSV of ~200,000

hr⁻¹, and thus is many-fold higher than that used in the Ford test [1]. However, the higher flow enabled the vapor transport of volatile components to be greatly accentuated by virtue of the law of mass action, so as to ease the burden of subsequent volatile component recovery and analysis. Since lesser amounts of volatile components would be transported and recovered using lower flow conditions, the present test is considered to be a very sensitive method for determining the component volatility.

[0030] After the exposure, each sample was digested with concentrated aqueous HF, and analyzed by ICPES for tungsten and vanadium. The detection limits are 2.5 μ g for each V and W per gram (ppm) of capture material (e.g., alumina).

[0031] The results, presented as an average of 4 separate runs, are given in Table 1.

[0032] Table 1

Description	V (ppm)	W (ppm)
Starting Catalyst (DT-52™ with 2wt% vanadia)	10075	72675
Recovered Catalyst	10425	71975
Starting Capture Material (gamma alumina)	0	0
Recovered Capture Material	19	4525
Mass Balance. %	104	105
Std. Dev., %	4.5	5.4

[0033] It can be seen that the loss of V to the vapor phase is minimal but measurable, while the loss of W is appreciable since the recovered capture alumina material contained roughly 0.45 wt% W. Also, it can be seen that the mass balances for each are essentially 100%, since the measured average mass balances are within one standard deviation of the theoretical value.

[0034] **Example 2- Demonstration of stability of capture material.**

[0035] This example demonstrates that once the tungsta and vanadia are present on the surface of the highly stable, high surface area capture material (in this case, gamma alumina) they are not volatile under the test conditions, even at exceedingly high exposure temperatures. Thus, a gamma alumina sample (Alfa Aesar, aluminum oxide, gamma, catalyst support, high surface area, bimodal) was loaded with 47013 ppm W and 11200 ppm V (by deposition from alkaline monoethanolamine solution) and subsequently calcined at 600°C for 6 hr in air. This catalyst sample was positioned upstream of a downstream

capture material (un-doped gamma alumina) as in Example 1. Separate tests demonstrated that the surface area of the W and V-doped alumina was 191 m²/g after exposure at 750°C for 16 hr in an atmosphere that contained 10% H₂O, which demonstrates the high stability of the material. The catalyst and capture materials were then exposed to the reactant stream at 750°C for 1 hr with a gas flow rate of 65L/hr, and the spent samples were recovered and analyzed. There was no measurable amount of W or V on the capture bed, so that these oxides exhibit no volatility when supported on high surface area, highly stable alumina support. In summary Examples 1 and 2 demonstrate two important discoveries, namely that both V and W, when supported on a titania of low thermal stability, exhibit measurable volatility at 750°C, but V and W do not exhibit measurable volatility at the same temperature when supported on high stability, high surface area alumina.

[0036] Examples 3 through 7: Evaluation of various vanadia-based of catalyst materials.

[0037] The discoveries described in Examples 1 and 2 presented a method of the presently disclosed and claimed inventive concepts as a means to investigate the volatility of various catalyst components such as V and W at the laboratory scale. Thus, in the following examples, SCR catalysts that contain tungsta and vanadia supported on various titania-based supports, were screened for the volatility of the catalyst components. The materials in these Examples all contained 2 wt% V₂O₅, deposited from alkaline monoethanolamine solution, and the V-doped materials were calcined at 600°C for 6 hr in air to remove the water and organic components. The DT-58™ base material is a commercially available titania-based SCR catalyst support available from MIC. DT-58™ comprises 81% TiO₂, 9% WO₃ and 10% SiO₂ and has a BET surface area of about 90-110 m²/g. The samples in Examples 5 and 6, labeled MC-X5v1 and MC-X5v2, are developmental SCR catalyst supports as described hereinbelow in reference to "Stabilized Anatase Titania" and in U.S. Patent Application Serial No. 12/533,414 which is expressly incorporated herein by reference in its entirety. The composition of the MC-X5v1 support is 90% TiO₂, 4% SiO₂ and 6% WO₃, while the composition of the MC-X5v2 support is 88 %TiO₂, 8% SiO₂ and 4% WO₃.

[0038] Table 2

Ex.	Material [a]	Condition	Surface Area (m ² /g)	Pore Volume (cm ³ /g)	V (ppm)	W (ppm)	TFm
4a	DT-58™	670C, 64 h, 10% H ₂ O, 6 L/h	55.7	0.25	N/A	N/A	N/A
4	DT-58™ (b)	750C, 4 h, 5% H ₂ O, 32.5 L/h	43.0	0.28	42.5	7150	5.6
5	MC-X5v1 (b)	750C, 4 h, 5% H ₂ O, 32.5 L/h	38.7	0.25	1.7	753.5	3.1
3	DT-52™ (b)	750C, 4 h, 5% H ₂ O, 32.5 L/h	8.3	0.05	82.6	8,050	9.3
6	MC-X5v2 (b)	750C, 4 h, 5% H ₂ O, 32.5 L/h	59.6	0.29	0.0	190.5	3.0

(a) = 2 wt% V₂O₅

(b) = average of multiple runs

[0039] It was first desired to determine a set of conditions that would be representative of the real-life exposure of a catalyst. As described above, Ford researchers[1] have developed an accelerated aging protocol for SCR catalysts that are positioned downstream of a diesel particulate filter (DPF), that simulates on-road exposure for 120,000 miles. The accelerated aging test involves exposing the catalyst to a reactant gas stream that includes water (5 vol%), for a time of 64 hr at 670°C at GHSV = 30,000 hr⁻¹. These exposure conditions represent the extremely harsh conditions that would occur as the result of the high temperatures created during soot-combustion during regeneration of the DPF, and are not normally encountered in traditional SCR applications.

[0040] Thus, the DT-58™ and MC-X5v1 catalysts were exposed for 64 hr at 670°C in an atmosphere that contained 10 vol% H₂O, and the surface area and pore volumes of the aged catalysts was determined as shown in Table 2. The same starting catalyst materials were also utilized as the catalyst samples in a dual bed configuration upstream of a downstream gamma alumina capture bed, and were treated at 750°C as described in Examples 1 and 2, above, only the exposure time was 4 hr and the reactant gas flow was 32.5 L/hr (equivalent to GHSV ~ 100,000 hr⁻¹). The gas flow rate under these conditions is still higher than in the Ford test[1], but is more closely representative of that test.

[0041] The results in Table 2 show that the surface areas for the catalysts exposed in the test method of the presently disclosed and claimed inventive concepts at 750°C for 4 hr were slightly lower than for the same materials aged at 670°C for 64 hr. Thus, if the surface area is used as a measure of the extent to which the catalyst samples have aged (and hence severity), then the

conditions in the former test are slightly more severe than in the Ford test[1]. Thus, it is concluded that the test of the presently disclosed and claimed inventive concepts, when conducted at 750°C for 4 hr with a reactant gas flowing at 32.5 L/hr, is a good first approximation to the real-life exposure of the catalyst over the duration of 120,000 miles of on-road use.

[0042] Each of the catalysts listed in Table 2 were evaluated multiple times and average results are shown therein. The results in Table 2 demonstrate that the catalysts show varying degrees of loss of V and W. Also shown in Table 2 are the Total Fractional Monolayer (Tf_m) values for each material, where the surface area is that of the aged samples. Shown in Figures 3 and 4 are the correlations found between the amount of V and W lost from the catalyst samples and the Tf_m values. The data and graphs show that when the Tf_m values are equal to or less than about 3, the vanadia and tungsta catalyst components exhibit low volatility and are substantially retained on the catalyst material, while when the Tf_m values are greater than 3, the catalyst components exhibit much higher volatility and are lost from the catalyst material (but are captured on the capture material). Of course, even with the relatively high levels of V and W that are lost from the DT-52™ catalyst material, when these volatile oxides and hydroxides are captured on the capture bed material, e.g., alumina, the Tf_m for that material is much less than 1 after the exposure. These good correlations can be used to predict the behavior of unknown materials that might be useful as catalyst and capture materials. Thus, vanadia-based titania materials that have high aged surface areas and low amounts of added catalyst components such as SiO_2 and WO_3 will demonstrate little to no loss of W and V to the vapor phase under these harsh exposure conditions, and hence are attractive catalyst materials. Such support materials will also be good capture materials in the event that less stable, high Tf_m catalyst materials such as vanadia on DT-52™ are positioned upstream or in intimate mixture with the more stable, low Tf_m materials.

[0043] Without being bound by theory, it is believed that the reason for the low volatility of the minority-phase components on alumina or stabilized titania (or other materials contemplated herein) is that when the minority-phase components are present at low fractional coverage (low f_m), they interact strongly chemically with the majority-phase support, and this favorable

interaction energy in turn lowers the equilibrium constants involved in the vaporization of the supported components.

[0044] Stabilized Anatase Titania.

[0045] In preferred embodiments of the presently disclosed and claimed inventive concepts, the material which comprises the majority phase catalyst material and/or the capture bed material used herein is an anatase titania (for example as described in U.S. Serial No. 12/533,414), wherein the anatase titania is stabilized by a silica provided in a low molecular weight form and/or small nanoparticle form. Further, the minority phase disposed on the anatase titania preferably comprises vanadia (and optionally tungsta), for vanadia-based selective catalytic reduction of DeNOx from lean-burn (diesel) engines.

[0046] The actual specific composition of the silica-titania or silica-titania-tungsta catalyst support material (majority phase) and/or capture bed material may be dictated by the requirements of the specific catalytic application. In one preferred composition, the material comprises a silica-stabilized titania material which comprises particles which contain $\geq 90\%$ dry weight of TiO_2 and ≤ 10 wt% SiO_2 . In another preferred composition, the material comprises a silica stabilized titania-tungsta material with $\geq 85\%$ dry weight titania, 3%-10% dry weight of SiO_2 , and 3%-10% dry weight of WO_3 . Alternatively, in one embodiment where the application requires particularly good thermal stability, the material comprises $\geq 85\%$ dry weight of TiO_2 , 5.0%-9.0% dry weight of SiO_2 , and 3.0%-7.0% dry weight of WO_3 . More particularly, the material may comprise 87%-89% dry weight of TiO_2 , 7%-9% dry weight of SiO_2 , and 3%-5% dry weight of WO_3 . In one preferred embodiment the material comprises about 88% ($\pm 0.5\%$) dry weight TiO_2 , about 8% ($\pm 0.5\%$) dry weight SiO_2 , and about 4% ($\pm 0.5\%$) dry weight WO_3 . In one embodiment, the weight % of WO_3 is less than the weight of % of SiO_2 . In one embodiment, the material has a fresh surface area of at least 80 m^2/gm , and more preferably at least 100 m^2/gm .

[0047] In another embodiment where the application requires particularly good catalytic activity or capture of volatiles, the material comprises $\geq 85\%$ dry weight of TiO_2 , 3.0%-8.0% dry weight of SiO_2 , and 4.0%-9.0% dry weight of WO_3 . More particularly, this active material

comprises $\geq 87\%$ dry weight of TiO_3 , 3%-6% dry weight of SiO_2 , and 4%-8% dry weight of WO_3 . In one preferred embodiment the material comprises about 90% ($\pm 0.5\%$) dry weight TiO_2 , about 4% ($\pm 0.5\%$) dry weight SiO_2 , and about 6% ($\pm 0.5\%$) dry weight WO_3 . In one embodiment, the weight % of WO_3 is greater than the weight % of SiO_2 .

[0048] In an embodiment of the presently disclosed and claimed inventive concepts, the TiO_2 component of the material used herein substantially comprises a surface area $< 400 \text{ m}^2/\text{g}$ and a pore volume $< 0.40 \text{ cm}^3/\text{g}$.

[0049] In one embodiment, the material is produced by mixing a titania slurry and a silica component at a temperature $< 80^\circ\text{C}$ and at a pH < 8.5 . Alternatively, the titania slurry and silica component used herein may be mixed at a temperature $< 70^\circ\text{C}$ and at a pH < 7.0 .

[0050] The vanadia catalyst of the presently disclosed and claimed inventive concepts may comprise the silica-stabilized titania or titania-tungsta catalyst support described herein upon which a quantity of vanadium oxide (V_2O_5) is disposed, wherein the V_2O_5 preferably comprises 0.5% to 1% to 2% to 3% to 4% to 5% of the dry weight thereof. The vanadia catalyst materials of the presently disclosed and claimed inventive concepts may be further treated by calcination (sintering) at a temperature $\geq 650^\circ\text{C}$ to increase their deNOx catalytic activity.

[0051] The emission system of the presently disclosed and claimed inventive concepts with may be used with a diesel particulate filter (DPF) upstream of the engine or downstream of the engine.

[0052] Preferably, most of the silica particles in the silica-stabilized titania material have diameters $< 5 \text{ nm}$, and more preferably $< 4 \text{ nm}$ and more preferably $< 3 \text{ nm}$, and still more preferably $< 2 \text{ nm}$, and/or comprise low molecular weights (e.g., $\text{MW} < 100,000$), whether or not the particles do, or do not, have V_2O_5 deposited thereon.

[0053] Where the silica titania material also has V_2O_5 deposited thereon the V_2O_5 preferably comprise from 0.5% to 1% to 2% to 3.0% to 4% to 5% of the dry weight of the material.

[0054] Distribution of the WO_3 and SiO_2 species on the surface of the titania material also plays a role in the optimization of DeNOx activity of the vanadia catalysts. Thus, when the catalysts are freshly prepared, that is,

when the added silica and tungsta oxides are first deposited and before high temperature treatment, the fractional monolayer coverage is preferably about 1.0 or less.

[0055] The SiO₂ may be present at a fractional monolayer value of less than 1.0 before the material is sintered. The small nanoparticle form of the SiO₂ may comprise a diameter of <5nm. The low molecular weight form of the SiO₂ may comprise a MW of < 100,000. The SiO₂ may comprise silicon atoms which are substantially (e.g., >50%) in the Q³, Q², Q¹ and Q⁰ coordination environments. The SiO₂ may comprise patches which are substantially ≤5 nm deep after redistribution as seen by scanning electron microscopy or by transmission electron microscopy. The TiO₂ used may optionally not be prepared in the presence of urea.

[0056] In another aspect, the invention may be a vanadia catalyst comprising a silica-stabilized titania material as described herein which comprises V₂O₅ disposed thereon. The vanadia catalyst may comprise, for example, 0.5% to 1 % to 2% to 3% to 4% to 5% dry weight of V₂O₅ (or more preferably 1.0% to 2% to 3%). The V₂O₅ may be present at a fractional monolayer value of less than 1.0 before sintering. The vanadia catalyst may be sintered at ≥ 650°C for example. In another aspect, the system of the presently disclosed and claimed inventive concepts may be a diesel selective catalytic reduction catalyst system comprising the vanadia catalyst and capture bed material as described herein. In another aspect, a diesel engine exhaust treatment system may further comprise a diesel particulate filter, and wherein the present catalytic capture bed device is positioned upstream of or downstream of the diesel particulate filter.

[0057] In another one of its aspects, the presently disclosed and claimed inventive concepts comprises a method of catalyzing the conversion of nitrogen oxides to N₂ gas, comprising exposing engine emissions comprising NO_x to the vanadia catalyst as described herein with an added reductant to produce N₂ and H₂O. The reductant may be for example NH₃ and/or urea. In the method the vanadia catalyst may comprise 0.5%-5% (or more preferably 1.0% to 3%) dry weight of V₂O₅, for example. The engine emissions may be passed through a diesel particulate filter before or after being exposed to the vanadia catalyst wherein the emissions are then passed through the capture bed material.

[0058] As noted above, the stabilization of titania material with silica preferably involves treatment of the titania with silica in a low molecular weight form and/or small nanoparticle form, such as tetra(alkyl)ammonium silicate (e.g., tetramethylammonium silicate) or tetraethylorthosilicate (TEOS). Other examples of low molecular weight and/or small nanoparticle silica precursors which may be used in the present invention include, but are not limited to aqueous solutions of silicon halides (i.e., anhydrous SiX_4 , where X= F, Cl, Br, or I), silicon alkoxides (i.e., Si(OR)_4 , where R=methyl, ethyl, isopropyl, propyl, butyl, iso-butyl, see-butyl, tert-butyl, pentyls, hexyls, octyls, nonyls, decyls, undecyls, and dodecyls, for example), other silicon-organic compounds such as hexamethyldisilazane, fluoro-silicic acid salts such as ammonium hexafluorosilicate $[(\text{NH}_4)_2\text{SiF}_6]$, quaternary ammonium silicate solutions (e.g., $(\text{NR}_4)_n$, (SiO_2) , where R=H, or alkyls such as listed above, and when $n=0.1-2$, for example), aqueous sodium and potassium silicate solutions (Na_2SiO_3 , K_2SiO_3 , and MSiO_3 wherein M is Na or K in varying amounts in ratio to Si), silicic acid (Si(OH)_4) generated by ion exchange of any of the cationic forms of silica listed herein using an acidic ion-exchange resin (e.g., ion-exchange of the alkali-silicate solutions or quaternary ammonium silicate solutions). In preferred embodiments, the titania used herein has not been prepared in the presence of urea.

[0059] The catalyst support material and/or capture bed material may be produced by providing a slurry comprising TiO_2 , combining the TiO_2 slurry with (1) a silica precursor solution comprising SiO_2 substantially in a low molecular weight form and/or SiO_2 comprising small nanoparticles and with (2) WO_3 to form a TiO_2 - WO_3 - SiO_2 mixture, wherein the silica precursor solution is combined with the TiO_2 slurry before, after, or while the WO_3 is combined with the TiO_2 slurry, and then washing and sintering the TiO_2 - WO_3 - SiO_2 mixture to form a silica-stabilized titania support material. In the method the silica-stabilized titania support material may comprise, for example, 86%-94% dry weight of TiO_2 , 3%-9% dry weight of a SiO_2 , and 3%-7% dry weight of WO_3 , and the titania support material may primarily comprise a surface area of at least 80 m^2/gm before sintering. The TiO_2 of the slurry may comprise, for example, preformed titanium hydroxide, titanium oxy-hydroxide or titanium dioxide particles. Optionally, the TiO_2 of the slurry is not produced in the presence of urea.

[0060] The silica precursor solution may comprise a tetra(alkyl)ammonium silicate solution or silicic acid. The SiO₂ may substantially comprise patches which are ≤5 nm in depth after redistribution as seen by scanning electron microscopy or by transmission electron microscopy. The method may further comprise combining the TiO₂-WO₃-SiO₂ mixture with V₂O₅ to form a vanadia catalyst. The vanadia catalyst thus formed may comprise, for example, 0.5% to 1% to 2% to 3% to 4% to 5% dry weight of V₂O₅. The V₂O₅ thereof may be present at a fractional monolayer value of less than 1.0 before sintering. The vanadia catalyst may be sintered at ≥650°C for example.

[0061] Alternatively, the silica-stabilized titania material may be produced by providing a TiO₂ slurry comprising TiO₂ particles, providing a particulate silica source, combining the TiO₂ slurry with the particulate silica source to form a TiO₂-SiO₂ mixture, and adjusting the TiO₂-SiO₂ mixture to a pH <8.5 and a temperature <80°C wherein the particulate silica source is dissolved and reprecipitated on the TiO₂ particles to form the silica-stabilized titania material. The method may further comprise the step of combining the silica-stabilized titania material with WO₃ to form a silica-stabilized titania tungsten material. The method may further comprise washing and sintering the silica-stabilized titania tungsten material. The silica-stabilized titania tungsten material may comprise, for example, 86%-94% dry weight of TiO₂, 3%-9% dry weight of a SiO₂, and 3%-7% dry weight of WO₃, and the titania material may primarily comprise a surface area of at least 80 m²/gm before sintering. The TiO₂ particles of the TiO₂ slurry may comprise, for example, preformed titanium hydroxide, titanium oxy-hydroxide or titanium dioxide particles. The SiO₂ of the TiO₂-SiO₂ mixture, after dissolving, may comprise silicon atoms which are substantially (e.g., > 50%) in the Q³, Q², Q¹ and Q⁰ coordination environments. The SiO₂ on the TiO₂ particles of the method may substantially comprise patches which are ≤5 nm in depth after redistribution of the SiO₂ as seen by scanning electron microscopy or by transmission electron microscopy. The method may further comprise combining the TiO₂-WO₃-SiO₂ mixture with V₂O₅ to form a vanadia catalyst. In the method, the vanadia catalyst may comprise, for example, 0.5% to 1% to 2% to 3% dry weight of V₂O₅. The V₂O₅ of the

vanadia catalyst may be present at a fractional monolayer value of less than 1.0 before sintering, and the vanadia catalyst may be sintered at $\geq 650^{\circ}\text{C}$.

[0062] As contemplated herein, in one embodiment, the presently disclosed and claimed inventive concepts comprises a selective catalytic reduction catalyst system for treating diesel exhaust gas containing nitrogen oxides and diesel soot particulates. The system comprises a catalyst material and a capture material. The catalyst material comprises a majority phase which may comprise a titania-based support material, and a minority phase comprising a catalyst component comprising at least one oxide of vanadium, silicon, tungsten, molybdenum, iron, cerium, phosphorous, copper or manganese. The capture material comprises a majority phase for capturing a minority phase comprising volatile oxides or hydroxides originating from the catalyst material, wherein the minority phase of the capture material maintains a total fractional monolayer coverage on the majority phase of the capture material of about 5 or less. The capture material is positioned in a mixture with the catalyst material, or is located downstream of the catalyst material or is positioned in a mixture with the catalyst material and is located downstream of the catalyst material. The minority phase of the capture material in this embodiment may maintain a total fractional monolayer coverage of 5 or less on the majority phase of the capture material when exposed to conditions of 750°C for 4 hours at a gas-hourly space velocity of $100,000\text{ hr}^{-1}$ and 5 vol% water. The minority phase catalyst components of the catalyst material may maintain a total fractional monolayer value of 5 or less on the majority phase after exposure to conditions of 750°C for 4 hours at a gas-hourly space velocity of $100,000\text{ hr}^{-1}$ and 5 vol% water. The capture material is preferably able to remove substantially all volatile oxides and hydroxides originating from the catalyst material. The majority phase of the capture material may primarily comprise at least one of aluminas, stabilized aluminas, silicas, silica-aluminas, amorphous silicas, titanias, silica-stabilized titanias, zeolites or molecular sieves or combinations thereof. Where the majority phase is a stabilized alumina, the stabilized aluminas may be stabilized by lanthanum or other lanthanides. The majority phase of the capture material and the majority phase of the catalyst material may comprise titania stabilized with silica.

[0063] In another embodiment the presently disclosed and claimed inventive concepts comprise a diesel engine exhaust treatment system

comprising a selective catalytic reduction catalyst system and a diesel particulate filter for treating diesel exhaust gas containing nitrogen oxides and diesel soot particulates. The catalyst system comprises a catalyst material and a capture material. The catalyst material comprises a majority phase which may comprise a titania-based support material, and a minority phase comprising a catalyst component comprising at least one oxide of vanadium, silicon, tungsten, molybdenum, iron, cerium, phosphorous, copper or manganese. The capture material may comprise a majority phase for capturing a minority phase comprising volatile oxides or hydroxides originating from the catalyst material, wherein the minority phase of the capture material maintains a total fractional monolayer coverage on the majority phase of the capture material of about 5 or less. The capture material may be positioned in a mixture with the catalyst material, or may be located downstream of the catalyst material, or may be positioned in a mixture with the catalyst material and located downstream of the catalyst material. The selective catalytic reduction catalyst system may be positioned upstream of or downstream of the diesel particulate filter. The minority phase of the capture material in this embodiment may maintain a total fractional monolayer coverage of 5 or less on the majority phase of the capture material when exposed to conditions of 750°C for 4 hours at a gas-hourly space velocity of 100,000 hr⁻¹ and 5 vol% water. The minority phase catalyst components of the catalyst material may maintain a total fractional monolayer value of 5 or less on the majority phase after exposure to conditions of 750°C for 4 hours at a gas-hourly space velocity of 100,000 hr⁻¹ and 5 vol% water. The capture material is preferably able to remove substantially all volatile oxides and hydroxides originating from the catalyst material. The majority phase of the capture material may primarily comprise at least one of aluminas, stabilized aluminas, silicas, silica-aluminas, amorphous silicas, titanias, silica-stabilized titanias, zeolites or molecular sieves or combinations thereof. Where the majority phase is a stabilized alumina, the stabilized aluminas may be stabilized by lanthanum or other lanthanides. The majority phase of the capture material and the majority phase of the catalyst material may comprise titania stabilized with silica.

[0064] In another embodiment, the presently disclosed and claimed inventive concepts comprise a method of treating diesel exhaust gas comprising the steps of providing a selective catalytic reduction catalyst system, with or

without a diesel particulate filter, and exposing the diesel exhaust gas to the selective catalytic reduction catalyst system wherein the capture material removes substantially all volatile oxides and hydroxides originating from the catalyst material from the diesel exhaust gas. In this method the selective catalytic reduction catalyst system comprises a catalyst material and a capture material. The catalyst material comprises a majority phase which may comprise a titania-based support material, and a minority phase comprising a catalyst component comprising at least one oxide of vanadium, silicon, tungsten, molybdenum, iron, cerium, phosphorous, copper or manganese. The capture material comprises a majority phase for capturing a minority phase comprising volatile oxides or hydroxides originating from the catalyst material, wherein the minority phase of the capture material maintains a total fractional monolayer coverage on the majority phase of the capture material of about 5 or less. The capture material is positioned in a mixture with the catalyst material, or is located downstream of the catalyst material or is positioned in a mixture with the catalyst material and is located downstream of the catalyst material.

[0065] The minority phase of the capture material in this embodiment may maintain a total fractional monolayer coverage of 5 or less on the majority phase of the capture material when exposed to conditions of 750°C for 4 hours at a gas-hourly space velocity of 100,000 hr⁻¹ and 5 vol% water. The minority phase catalyst components of the catalyst material may maintain a total fractional monolayer value of 5 or less on the majority phase after exposure to conditions of 750°C for 4 hours at a gas-hourly space velocity of 100,000 hr⁻¹ and 5 vol% water. The capture material is preferably able to remove substantially all volatile oxides and hydroxides originating from the catalyst material. The majority phase of the capture material may primarily comprise at least one of aluminas, stabilized aluminas, silicas, silica-aluminas, amorphous silicas, titanias, silica-stabilized titanias, zeolites or molecular sieves or combinations thereof. Where the majority phase is a stabilized alumina, the stabilized aluminas may be stabilized by lanthanum or other lanthanides. The majority phase of the capture material and the majority phase of the catalyst material may comprise titania stabilized with silica.

[0066] While the presently disclosed and claimed inventive concepts been described in connection with certain preferred embodiments and examples herein so that aspects thereof may be more fully understood and appreciated, it

is not intended to limit the presently disclosed and claimed inventive concepts to these particular embodiments and examples. Thus, the present examples, which include preferred embodiments serve to illustrate the practice of this invention, it being understood that the particulars shown are by way of example and for purposes of illustrative discussion of preferred embodiments of the presently disclosed and claimed inventive concepts only and are presented in the cause of providing what is believed to be the most useful and readily understood description of formulation procedures as well as of the principles and conceptual aspects of the presently disclosed and claimed inventive concepts.

[0067] Therefore, although the presently disclosed and claimed inventive concepts and their advantages have been described in detail, it should be understood that various changes, substitutions and alterations can be made herein without departing from the spirit and scope of the presently disclosed and claimed inventive concepts as defined by the appended claims. Moreover, the scope of the present application is not intended to be limited to the particular embodiments of the machines, processes, items of manufacture, compositions of matter, means, methods and steps described in the specification. As one of ordinary skill in the art will readily appreciate from the present disclosure, machines, processes, items of manufacture, compositions of matter, means, methods, or steps, presently existing or later to be developed that perform substantially the same function or achieve substantially the same result as the corresponding embodiments described herein may be utilized according to the presently disclosed and claimed inventive concepts. Accordingly, the appended claims are intended to include within their scope such machines, processes, items of manufacture, compositions of matter, means, methods, or steps.

[0068] Each of the patents, published patent applications, references and articles cited herein is hereby expressly incorporated herein by reference in its entirety.

[0069] CITED REFERENCES

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- [0071]** 2. Weckhuysen, B.M., and Keller, D.E., "Chemistry, Spectroscopy and the Role of Supported Vanadium Oxides in Heterogeneous Catalysis", *Catalysis Today* 78:25-46, 2003.
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What is claimed is:

1. A selective catalytic reduction catalyst system for treating diesel exhaust gases which contain nitrogen oxides and diesel soot particulates, comprising:
 - a catalyst material comprising:
 - a majority phase comprising a titania-based support material, and
 - a minority phase comprising a catalyst component comprising at least one oxide of vanadium, silicon, tungsten, molybdenum, iron, cerium, phosphorous, copper or manganese; and
 - a capture material comprising a majority phase for capturing a minority phase comprising volatile oxides or hydroxides originating from the catalyst material, wherein the minority phase of the capture material maintains a total fractional monolayer coverage on the majority phase of the capture material of about 5 or less; andwherein the capture material is positioned in a mixture with the catalyst material, or is located downstream of the catalyst material.
2. The selective catalytic reduction catalyst system of claim 1 wherein the minority phase of the capture material maintains the total fractional monolayer coverage of 5 or less on the majority phase of the capture material when exposed to conditions of 750°C for 4 hours at a gas-hourly space velocity of 100,000 hr⁻¹ and 5 vol% water.
3. The selective catalytic reduction catalyst system of claim 1 wherein the minority phase catalyst components of the catalyst material maintain a total fractional monolayer value of 5 or less on the majority phase after exposure to conditions of 750°C for 4 hours at a gas-hourly space velocity of 100,000 hr⁻¹ and 5 vol% water.
4. The selective catalytic reduction catalyst system of claim 1 wherein the capture material is able to remove substantially all volatile oxides and hydroxides originating from the catalyst material.
5. The selective catalytic reduction catalyst system of claim 1 wherein the majority phase of the capture material primarily comprises at least one of aluminas, stabilized aluminas, silicas, silica-aluminas, amorphous silicas,

titanias, silica-stabilized titanias, zeolites or molecular sieves or combinations thereof.

6. The selective catalytic reduction catalyst system of claim 5 wherein the majority phase of the capture material primarily comprises stabilized aluminas which are stabilized by lanthanum or other lanthanides.

7. The selective catalytic reduction catalyst system of claim 1 wherein the majority phase of the capture material and the majority phase of the catalyst material comprise titania stabilized with silica.

8. A diesel engine exhaust treatment system, comprising:
- a selective catalytic reduction catalyst system for treating diesel exhaust gases which contain nitrogen oxides and diesel soot particulates, comprising:
 - a catalyst material comprising:
 - a majority phase comprising a titania-based support material, and
 - a minority phase comprising a catalyst component comprising at least one oxide of vanadium, silicon, tungsten, molybdenum, iron, cerium, phosphorous, copper or manganese; and
 - a capture material comprising a majority phase for capturing a minority phase comprising volatile oxides or hydroxides originating from the catalyst material, wherein the minority phase of the capture material maintains a total fractional monolayer coverage on the majority phase of the capture material of about 5 or less; and
 - wherein the capture material is positioned in a mixture with the catalyst material, or is located downstream of the catalyst material; and
 - a diesel particulate filter, and wherein the selective catalytic reduction catalyst system is positioned upstream of or downstream of the diesel particulate filter.

9. The diesel engine exhaust treatment system of claim 8 wherein the minority phase of the capture material of the selective catalytic reduction catalyst system maintains the total fractional monolayer coverage on the majority phase of the capture material of 5 or less when exposed to conditions of 750°C for 4 hours at a gas-hourly space velocity of 100,000 hr⁻¹ and 5 vol% water.

10. The diesel engine exhaust treatment system of claim 8 wherein the minority-phase catalyst components of the catalyst material of the selective catalytic reduction catalyst system maintain a total fractional monolayer value of 5 or less on the majority phase after exposure to conditions of 750°C for 4 hours at a gas-hourly space velocity of 100,000 hr⁻¹ and 5 vol% water.

11. The diesel engine exhaust treatment system of claim 8 wherein the capture material of the selective catalytic reduction catalyst system is able to remove substantially all volatile oxides and hydroxides originating from the catalyst material.

12. The diesel engine exhaust treatment system of claim 8 wherein the majority phase of the capture material of the selective catalytic reduction catalyst system primarily comprises at least one of aluminas, stabilized aluminas, silicas, silica-aluminas, amorphous silicas, titanias, silica-stabilized titanias, zeolites or molecular sieves or combinations thereof.

13. The diesel engine exhaust treatment system of claim 12 wherein the majority phase of the capture material primarily comprises stabilized aluminas which are stabilized by lanthanum or other lanthanides.

14. The diesel engine exhaust treatment system of claim 8 wherein the majority phase of the capture material and the majority phase of the catalyst material of the selective catalytic reduction catalyst system comprise titania stabilized with silica.

15. A method of treating diesel exhaust gas, comprising:
providing a selective catalytic reduction catalyst system comprising:
a catalyst material comprising:

- a majority phase comprising a titania-based support material, and a minority phase comprising a catalyst component comprising at least one oxide of vanadium, silicon, tungsten, molybdenum, iron, cerium, phosphorous, copper or manganese; and
- a capture material comprising a majority phase for capturing a minority phase comprising volatile oxides or hydroxides originating from the catalyst material, wherein the minority phase of the capture material maintains a total fractional monolayer coverage on the majority phase of the capture material of about 5 or less; and
- wherein the capture material is positioned in a mixture with the catalyst material, or is located downstream of the catalyst material; and
- exposing the diesel exhaust gas to the selective catalytic reduction catalyst system wherein the capture material removes substantially all volatile oxides and hydroxides originating from the catalyst material from the diesel exhaust gas.

16. The method of claim 15 wherein the minority phase of the capture material of the selective catalytic reduction catalyst system maintains the total fractional monolayer coverage on the majority phase of the capture material of 5 or less when exposed to conditions of 750°C for 4 hours at a gas-hourly space velocity of 100,000 hr⁻¹ and 5 vol% water.

17. The method of claim 15 wherein the minority phase catalyst components of the catalyst material of the selective catalytic reduction catalyst system maintain a total fractional monolayer value of 5 or less on the majority phase after exposure to conditions of 750°C for 4 hours at a gas-hourly space velocity of 100,000 hr⁻¹ and 5 vol% water.

18. The method of claim 15 wherein the majority phase of the capture material of the selective catalytic reduction catalyst system primarily comprises at least one of aluminas, stabilized aluminas, silicas, silica-aluminas, amorphous silicas, titanias, silica-stabilized titanias, zeolites or molecular sieves or combinations thereof.

19. The method of claim 18 wherein the majority phase of the capture material primarily comprises stabilized aluminas which are stabilized by lanthanum or other lanthanides.

20. The method of claim 15 wherein the majority phase of the capture material and the majority phase of the catalyst material of the selective catalytic reduction catalyst system comprise titania stabilized with silica.

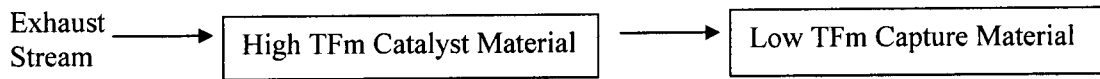


Figure 1

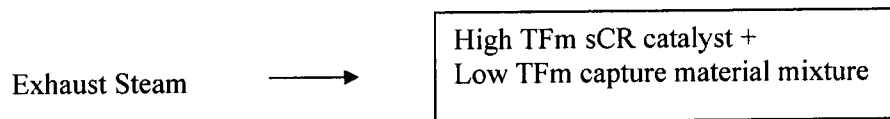


Figure 2

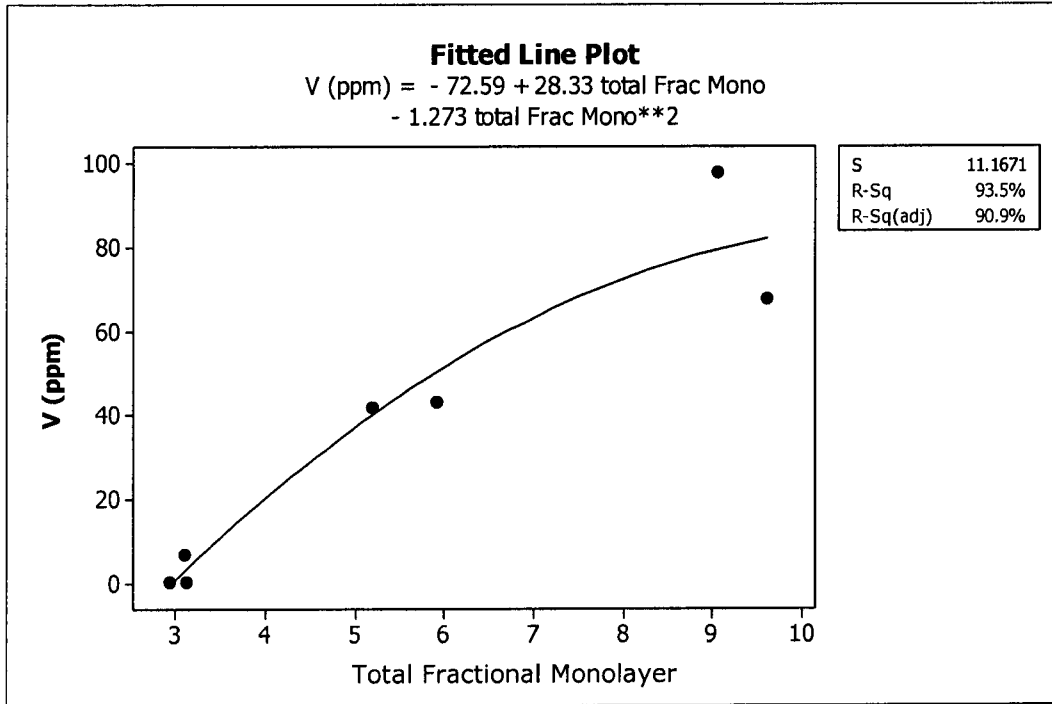


Figure 3

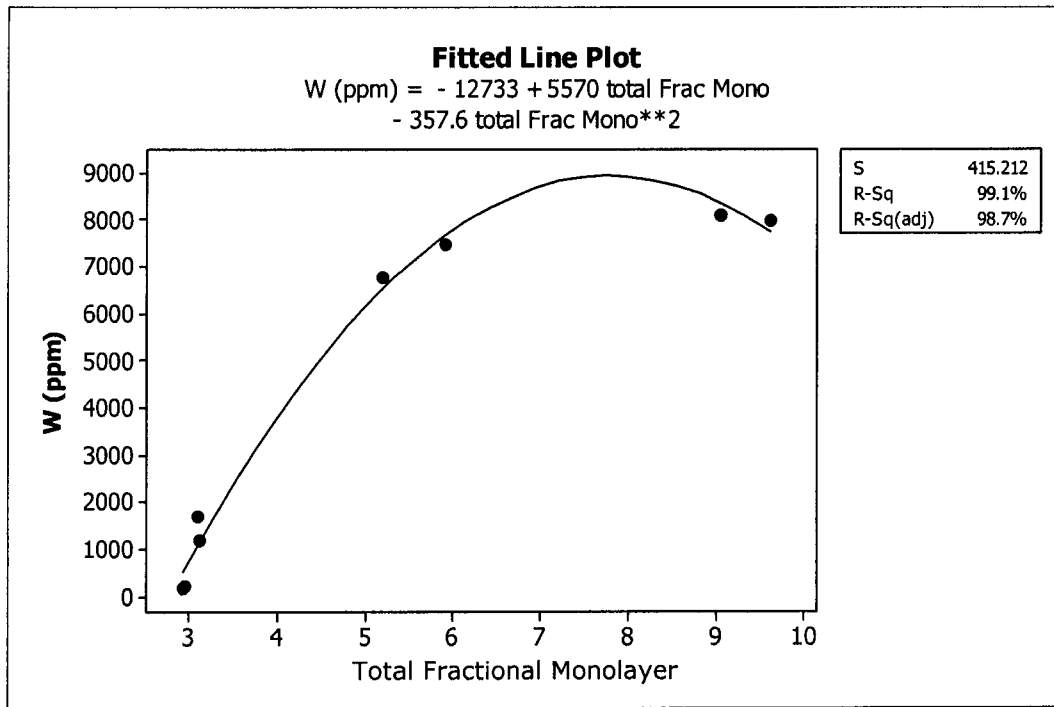


Figure 4